

**ATTACHMENT A**
Remarks

Claims 1-27 are pending in the present application. By this Amendment, Applicants have amended claims 14 and 27. Applicants respectfully submit that the present application is in condition for allowance based on the discussion which follows.

The abstract of disclosure was objected to for including a typographical error which, by this Amendment, Applicants have amended, thereby obviating the informality.

The specification was objected to for not including subject headings which, by this Amendment, Applicants have amended, in accordance with more conventional U.S. patent specification form.

Claims 14 and 27 were rejected under 35 U.S.C. § 112, second paragraph. By this Amendment, Applicants have amended claims 14 and 27, as suggested by the Examiner, to more clearly recite that the material selected is borosilicate glass.

Claims 1-2, 5-7 and 10-11 were rejected under 35 U.S.C. § 102(b) as being anticipated by Law, Jr. et al. (U.S. Patent No. 6,238,543) (hereinafter "Law").

Contrary to the Examiner's assertion, Law fails to teach or suggest the claimed method, which is clear from a thorough understanding and analysis of the structure of the polymer electrolyte membrane and other structures of the reactor used in Law.

The structure of the Law reactor is shown in Figures 1, 2A, 2B, 2C and 2D, and described, in particular, in lines 1 to 21 of column 4 of Law. Shown in Figure 2C of Law, the polymer electrolyte membrane reactor has an anode side, shown at the left hand side in Figure 2C, and a cathode side, shown at the right hand side of Figure 2C. The anode side and the cathode side are separated by a generally planar membrane electrode assembly (MEA). In Figure 2C, the MEA extends down the mid-line of

Figure 2C and is shown in cross-section as a thin member between the two Teflon[®] gaskets. Clearly, the plane in which the MEA lies is perpendicular to the plane of Figure 2C.

The MEA is shown in detail, in cross-section, in Figure 2D. As shown in Figure 2D, the MEA has a polymer electrolyte membrane (PEM) which is sandwiched between two parallel platinum electrodes, one of which is marked "Pt" in Figure 2D. At each side of the MEA, a respective platinum screen current collector is provided.

Turning back to Figure 2C, the MEA (and the Teflon[®] gaskets) are sandwiched between two graphite blocks (see lines 9 to 11 of column 4 of Law).

As made clear in lines 62 to 64 of column 1 of Law, the reaction between the two carboxylic acids takes place at the anode side of the reactor – that is to say the half of the reactor which is shown to the left of the MEA in Figure 2C. The carboxylic acids are introduced into the anode side of the reactor, and the products are removed from the anode side of the reactor, by a "flow field" (or channel network) which is shown in Figures 2C and 2A. Figure 2C shows one channel of the flow field which is labeled "gas in" and another channel of the flow field which is labeled "gas out." Although not shown in Figure 2C, the "gas in" channel and the "gas out" channel are connected by a network of channels – all of which lie to the left hand side of the MEA as shown in Figure 2C. It is clear from Figure 2A, that Figure 2C is a cross-section taken vertically through the mid point of Figure 2A. In Figure 2A, the "gas in" channel shown in Figure 2C is shown as a circular channel extending perpendicular to the page generally towards the top of Figure 2A and is labeled "gas inlet." The "gas out" channel shown in Figure 2C is seen in Figure 2A, as a circle, extending perpendicularly to the plane of the page, and is

labeled “gas outlet.” As shown in Figure 2A, the gas inlet and the gas outlet are connected by a network of channels which include four parallel vertical channels which extend between two horizontal parallel channels.

It should be emphasized that the gas inlet channel and the gas outlet channel, together with the connecting network of channels, are all located at the anode side of the reactor – that is to say at the left hand side of the MEA as shown in Figure 2C. There is a corresponding network of channels on the cathode side of the reactor, but there is clearly no direct communication between the anode channels and the cathode channels, because otherwise the reactor would not function.

It is also noted that the channels, which are referred to as a “flow field” in Law, are formed by machining in the respective graphite blocks, as discussed in lines 9 to 13 of column 4 of Law.

Having discussed the structure of the reactor described in Law, attention is now turned to the differences between claim 1 of the current application and the method of reacting carboxylic acids that is disclosed and described in Law.

Claim 1 of the current application requires:

“using electrodes to apply an electrical voltage between opposite ends of a channel containing a liquid.”

Additionally, claim 1 of the current application requires that:

“said reaction [that is to say the reaction between the carboxylic acid molecules] taking place in the liquid in the channel.”

In Law, the only channel in which carboxylic acids react is the channel network consisting of the gas inlet channel, the gas outlet channel and the connecting network of

channels which are formed in the graphite block of the anode side of the reactor.

However, in Law, there is no use of electrodes to apply an electrical voltage between opposite ends of this channel, as required by claim 1 of the current application. Instead, in the polymer electrolyte membrane reactor described in Law, the channels in the anode side of the reactor lie adjacent the anode of the MEA. However, there is no electrical voltage applied between opposite ends of the channel.

Moreover, one of ordinary skill in the art would not adapt or modify the reactor described in Law so as to apply an electrical voltage between opposite ends of the channels in the anode side of the reactor, because this would interfere with the proper functioning of the reactor. Accordingly, this is a clear distinction between claim 1 of the current application and Law. Further, claim 1 of the current application also requires that:

“said reaction taking place in the liquid in the channel and spaced from the electrodes.”

In contrast, in Law, Figure 1 indicates that the reaction between the carboxylic acids takes place at the anode surface. There is no suggestion whatsoever in Law that the reaction between the carboxylic acids may take place at a location spaced from the anode. Further, Law provides no motivation for one of ordinary skill in the art to modify the location of where the reaction is to take place.

Accordingly, it is respectfully submitted that claim 1 of the current application is not anticipated by Law.

Claim 2 of the current application requires that:

“the electrical voltage causes electro-osmotic movement of the liquid along the channel.”

In this regard, the Examiner alleges, on page 4 of the Office Action, that electro-osmotic movement of the liquid along the channels of Law is inherent. This is respectfully denied. For the following reasons, it is respectfully submitted that no electro-osmotic movement of liquid takes place in the channels of Law.

Firstly, in order to achieve electro-osmotic movement along a channel, it is necessary to have a voltage difference between the ends of the channel. For example, as shown in U.S. Patent No. 6,989,090 (also referred to by the Examiner), this is achieved by placing electrodes of opposite polarity at the ends of the channels. However, in the channels provided in the reactor of Law, there is no voltage difference between the ends of the channels, and so electro-osmotic flow cannot take place.

Another requirement for electro-osmotic flow is described in the right hand column of page 4R of the Review Article by S.J. Haswell, which was also referred to by the Examiner. Here, it is stated that:

“Firstly, in order to generate the EOF one must use a material which will yield negatively charged groups on the surface or walls of the channels when placed in contact with an appropriate liquid. Secondly, the liquid phase must dissociate to some extent in order to generate counter positive ions (notably H^+ ions). The combination of the negatively charged surface ... and the H^+ ions in solution will form a diffuse double layer This diffuse double layer acts as a parallel-plate electric capacitor ...”

In Law, the channels are formed by machining in the graphite blocks, and so the walls of the channels will be formed by a graphite. Additionally, the liquids are pure carboxylic acids (not solutions of carboxylic acids). It is respectfully submitted that this combination of graphite and pure carboxylic acid is unsuitable for the generation of an electro-osmotic force.

Based on the foregoing, Applicants respectfully submit that claims 1-2, 5-7 and 10-11 are not anticipated by Law under 35 U.S.C. § 102(b).

Claims 15, 18-20 and 23-24 were rejected under 35 U.S.C. § 102(b) as being anticipated by Law. Applicants respectfully traverse this rejection based on the following discussion.

Claim 15 of the current application requires:

“applying an electrical voltage to cause electro-osmotic movement of the liquid along the channel.”

However, for the reasons given above, this feature is not present in the reactor of Law. Moreover, electro-osmotic flow cannot take place in the reactor of Law, because the materials used are incompatible with electro-osmotic flow, and because the application of a voltage to cause electro-osmotic flow would interfere with the proper operation of the reactor of Law.

Based on the foregoing, Applicants respectfully submit that claims 15, 18-20 and 23-24 are not anticipated by Law.

Claims 1-14 were rejected under 35 U.S.C. § 103(a) as being unpatentable over Meresz et al. (U.S. Patent No. 4,006,065) (hereinafter “Meresz”) in view of Haswell et al. (U.S. Patent No. 6,989,090) (hereinafter “Haswell ‘090”) and Haswell, S.J., “Development and Operating Characteristics of Micro Flow Injection Analysis Systems Based on Electroosmotic Flow” (hereinafter “The Analyst article”).

Contrary to the § 103(a) rejection, Applicants respectfully submit that claims 1-14 are not obvious in view of the aforementioned references.

Meresz describes a reaction which is simply a variant of the classic Kolbe reaction. This is acknowledged by Meresz in line 35 of column 1. The Meresz reaction

is a variant because it uses an unsaturated carboxylic acid so as to form an unsaturated product.

On page 7 of the Office Action, the Examiner alleges that Meresz teaches that the reaction takes place at a location spaced from the electrodes. However, this is respectfully traversed. There is no teaching or suggestion in Meresz that the reaction takes place at a location spaced from the electrodes.

As discussed above, the Meresz reaction is simply a variant of the classic Kolbe reaction. The established theory relating to the Kolbe reaction is that the reaction between the carboxylate anions takes place at the electrode surface. In support of this we enclose a copy of a review article relating to the Kolbe reaction and entitled "Electrolytic Condensation of Carboxylic Acids" which appeared in volume 29 No. 3 of Russian Chemical Reviews in March 1960. This article makes it quite clear that the established wisdom with regard to the Kolbe type reaction is that the reaction takes place at the anode surface.

For example, in the right hand column of page 161 it is stated that:

"At present the term "Kolbe Synthesis" is understood to imply electrode condensation at the anode"

Further, in the right hand column of page 176, it is stated that:

"It can be assumed that the discharge of the carboxylic acid ion and its simultaneous interaction with the higher platinum oxides (formed at the electrode) is a necessary condition for the electrolytic condensation."

In view of this established wisdom, a person of ordinary skill in the art, reading Meresz, would have understood that the reaction between the carboxylic acid anions

takes place at the anode surface, not at a location spaced from the electrodes as required by claim 1 of the current application.

The Examiner alleges that it would be obvious to use the Meresz reaction in a micro-reactor of the type disclosed in Haswell '090, or in the Haswell Analyst article. However, this is respectfully traversed. In micro-reactors, reactions take place in the channels between the electrodes – not at the electrode surfaces. In view of this, the skilled person simply would not have attempted to use the Meresz reaction in a micro-reactor – because the skilled person would have understood that the Meresz reaction takes place at the anode surface, and so the skilled person would have reasoned that the reaction would not have taken place at all in the micro-reactor.

A thoroughly surprising aspect of the current invention is that a Kolbe type reaction can be performed in a micro-reactor and that the reaction takes place in the channels spaced from the electrodes. This would not have been anticipated by the skilled person in view of the cited documents.

Based on the foregoing, Applicants respectfully submit that claims 1-14 are not obvious under 35 U.S.C. § 103(a) in view of the aforementioned cited references.

In view of the foregoing, Applicants respectfully submit that the present application is in condition for allowance.

END REMARKS

ELECTROLYTIC CONDENSATION OF
CARBOXYLIC ACIDSG. E. Svadkovskaya and
S. A. Voitkevich

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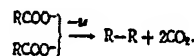
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INTRODUCTION

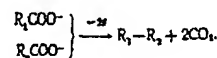
Investigations on electrolytic condensation processes and, in general, on organic electrosynthesis, were first carried out in 1849 by Kolbe who obtained ethane and carbon

dioxide at the anode during the electrolysis of an aqueous solution of potassium acetate (in the hope of separating the methyl radical)¹. Later it was found that the electrolysis products also contained ethylene and methyl acetate².

Kolbe believed that the substances formed at the anode were products of the partial oxidation of acetic acid with oxygen which was liberated at the anode, and therefore found it impossible to evaluate the reaction that he had discovered. Electrolytic condensation methods were used extensively in preparative organic chemistry as early as the second half of the 19th century. The fundamental nature of the Kolbe reaction became clear: a discharge of the carboxylic acid anions with subsequent decarboxylation and "fusion" of the radicals:



In 1855 Würtz^{3,4} first extended the use of the Kolbe synthesis by proving that it was possible to condense acids with varying numbers of carbon-atoms (the so-called "crossed" synthesis):



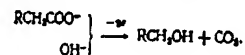
Brown and Walker⁵ found in 1891 that diesters of dicarboxylic acids could be synthesised by the electrolysis of salts of half-esters of aliphatic dicarboxylic acids:



Subsequent detailed investigations on the electrolytic synthesis of diethyl succinate from ethyl malonate⁶⁻⁸ confirmed that the Brown-Walker reaction was analogous to the electrolysis of salts of acetic acid and of other acids of the aliphatic series.

At present the term "Kolbe synthesis" is understood to imply electrocondensation at the anode not only of fatty acids but also of their various derivatives, such as the half-esters of dibasic acids. In this reaction water is usually used as solvent, but a series of publications quoted below describe the Kolbe synthesis in solvents such as alcohols, acetic acids, etc.

Side reactions usually occur during the formation of the desired compound, especially when the synthesis is carried out in an aqueous medium. This applies especially to the electrolysis of solutions of salts of carboxylic acids, as first shown by Hofer and Moest^{9,10}, when an alcohol containing one carbon-atom less than the original acid is also formed:



The formation of alcohol, according to the Hofer-Moest reaction, can become predominant in an alkaline medium and in the presence of certain anions; in this case the hydroxyl group can occupy not only the α -position in relation to the carboxyl group of the initial compound, but also the β - or even the γ -position¹¹⁻¹⁵. Frequently, especially at low current densities, saturated¹⁶⁻¹⁸ and unsaturated^{11,14,16,19-22} hydrocarbons* have been detected in the reaction products; these contain 2 carbon atoms less than the normal product:



* The use of this side reaction for the synthesis of muscone from (3-oxocyclopentadecyl-1)-acetic acid²³ is interesting.

It has been shown that esters^{2,16,24-28} and also aldehydes^{11,24,27} and ketones^{14,16} are formed as by-products during the Kolbe synthesis.

Despite these side reactions, the Kolbe condensation is a most convenient method for preparing various organic compounds. Recently numerous studies have been devoted to it and its practical applications. This review covers publications on anodic condensation reactions up to 1958*.

EFFECT OF EXPERIMENTAL CONDITIONS ON THE COURSE OF THE KOLBE SYNTHESIS

The yield of the principal electrocondensation product and the character of the side-reactions in the Kolbe synthesis are mainly determined by the reaction conditions (anodic current density, anode material, temperature, etc.). The effect of these conditions has been evaluated in detail for the electrolysis of aqueous solutions of alkali-metal acetates^{9,35-38} and the synthesis of diethyl succinate from the monoester of malonic acid.

It may be assumed that basic principles, established for these cases, apply generally to electrolytic condensation in aqueous media. The following account of the dependence of the electrolytic condensation on the conditions of electrolysis applies only to aqueous solutions.

1) Effect of Anode Material

In the electrolysis of acetates, the highest yields of ethane are obtained by using a smooth platinum or iridium anode³⁹; gold, nickel, and platinised platinum give negative results^{37,40-42}. Graphite anodes give moderate yields of ethane, and the quantity of ethane formed depends on the fineness of division of the graphite surface³⁷. Analogous effects were found on electrolysis of monoethyl malonate^{9,43}. Table 1 shows how the yield of ethane (from potassium acetate - Glasstone and Hickling³⁷) and of diethyl succinate (from the monoester of malonic acid - Hickling and Westwood⁹) depends on the anode material.

TABLE 1.

| Anode | Smooth Pt | Platinised Pt | Au | Ni | Retort carbon | Graphite | PbO ₂ |
|-------------------------------|-----------|---------------|----|----|---------------|----------|------------------|
| Yield of ethane, % | 80 | 3 | 0 | 0 | 21 | 4 | 0 |
| Yield of diethyl succinate, % | 74 | 0 | 12 | 0 | 26 | 35 | 0 |

TABLE 2.

| Current density, A/dm ² | Yield of ethane, % | Current density, A/dm ² | Yield of diethyl succinate, % |
|------------------------------------|--------------------|------------------------------------|-------------------------------|
| 0.5 | 52 | 1.0 | 40 |
| 2.5 | 71 | 5.0 | 48 |
| 5.0 | 77 | 10 | 68 |
| 25.0 | 89 | 30 | 76 |
| 50.0 | 89 | 30 | 74 |
| | | 100 | 69 |
| | | 200 | 66 |

* Some previously published reviews and monographs²⁸⁻³⁴ were consulted when compiling the present survey.

2) Effect of Current Density

In the electrolysis of solutions of acetic acid and its salts, increase of current density usually favours the synthesis⁴⁴. According to Glasstone and Hickling⁴⁵ the optimum current density for a smooth platinum anode in a solution containing equimolar quantities of acetic acid and potassium acetate is 25 A/dm². A further increase of current density does not increase the yield of ethane. A solution containing equimolar quantities of monoethyl malonate and of its potassium salt shows similar behaviour⁶ (Table 2).

3) Effect of Temperature

As a rule the yield of ethane on electrolysis of acetic acid-acetate solutions falls with increasing temperature^{36,46,47}. Thus for a 20% solution of potassium acetate, the yield of ethane decreases from 83% at 0° to 12% at 95°;^{36,46} for a 5% solution at 25 A/dm², the yield falls from 88.2% at 13° to 0.8% at 92°, and at 100° no ethane is formed⁴⁴. In other words, the adverse effect of raising the temperature becomes more obvious the more dilute the solution.

The yield of diethyl succinate on electrolysis of aqueous potassium ethyl malonate⁹ was 74% at 10° and only 22% at 90°; the current density was 50 A/dm² in both cases.

TABLE 3.

| Added salt | Yield, % | | Mobility of anions |
|------------------------------------|----------|-------------------|--------------------|
| | ethane | diethyl succinate | |
| Without additive | 71 | 70 | — |
| KF | 30 | 47 | 47 |
| KClO ₄ | 20 | 20 | 84 |
| KNO ₃ | 19 | 23 | 62 |
| KH ₂ PO ₄ | 17 | 22 | — |
| KCl | 5 | 19 | 65 |
| K ₂ SO ₄ | 3 | 11 | 69 |
| K ₄ Fe(CN) ₆ | 0 | 8 | 97 |

4) Effect of Foreign Anions

Foreign anions inhibit the principal anodic reaction and their presence leads to the formation of by-products. In particular, methanol is formed instead of ethane (Hofer-Moest reaction^{9,10,48}) when aqueous solutions of alkali-metal acetates are electrolysed in the presence of bicarbonates, sulphates, and perchlorates. The presence of foreign ions in the acetic acid solution may explain why some investigators did not succeed in preparing ethane by the Kolbe synthesis^{49,50}. Glasstone, Hickling, and Westwood^{6,37} have made a detailed study of the effect of such anions on the electrolysis of solutions of acetic acid and monoethyl malonate; the results are set out in Table 3.

The authors conclude that the inhibiting effect on the Kolbe synthesis increases with the mobility of the anion, and is considerably weaker in the electrolysis of solutions of salts of organic acids than in the electrolysis of aqueous solutions of the free acids.

5) Effect of Cations

The presence of certain metallic cations (Pb²⁺, Mn²⁺, Cu²⁺, Fe²⁺, and Co²⁺) in relatively small amounts strongly affects the course of the Kolbe reaction in solutions of alkali-metal acetates, reducing the yield of ethane and in

several cases completely suppressing its formation^{37,51}. For example, the addition of 0.001 M lead acetate to a solution containing acetic acid and potassium acetate in equimolar amounts reduces the yield of ethane from 70% to zero at current density 2.5 A/dm².⁵¹ The effect of the cations follows the order Pb²⁺ > Mn²⁺ > Cu²⁺ > Co²⁺ = Fe²⁺ and is especially marked at low current densities. Analogous results were obtained when the same cations were added to potassium ethyl malonate⁶.

However, the Kolbe synthesis is not inhibited by the presence of many other metal cations. Glasstone and Hickling³⁷ obtained ethane in satisfactory yields in the electrolysis of potassium, sodium, calcium, barium, and other acetate solutions (Table 4).

6) Effect of the pH of the Medium

The principal reaction product is usually formed when the pH of the electrolyte is low, its yield falling with increasing pH.^{54,55} According to Glasstone, Hickling, and Westwood^{6,49}, the optimum pH values lie within the limits 3-5 (Table 5); the yield of the product decreases rapidly in an alkaline electrolyte (at pH > 10).

7) Effect of Concentration

Increased concentration of carboxylic-acid anions usually leads to higher yields of the principal condensation product^{44,54,55}, the extent of side reactions increasing as the concentration is lowered. For example the yield of ethane decreases during the electrolysis of dilute potassium acetate solutions, more ethylene and methane being formed^{56,57}. Data^{6,37} illustrating the effect of concentration are set out in Table 6.

ELECTROLYTIC CONDENSATION IN NON-AQUEOUS SOLVENTS

The first experiments on the use of ethanol, propanol, and butanol as solvents in the electrolysis of acetates were carried out in 1886⁵⁸. The electrolysis of potassium acetate in glacial acetic acid was investigated in 1904.^{59,60} Further experiments were carried out on the electrolysis of acetates, propionates, and formates in anhydrous acetic, propionic, and formic acids¹⁹. Detailed studies proved^{40,44,55,61,62} that the anode material, current density, temperature, etc., have little effect on the yield of the product in the electrolysis of solutions of carboxylic-acid salts in non-aqueous solvents. Thus, the electrolysis of a potassium acetate solution in anhydrous acetic acid gives better yields of ethane when a gold or graphite anode is used instead of a platinum anode⁴⁴; the ratio C₂H₆ : H₂ equals 0.96 for a platinum anode, 0.95 for a gold anode, and 0.89 for a graphite anode; the addition of even 1% of water to the electrolyte, when a gold anode is used decreases the yield of ethane to 65%; the yield falls to 1.6% in the presence of 10% of water. Similar results were obtained in the electrolytic condensation of potassium propionate dissolved in anhydrous methanol⁴⁰.

Detailed studies of the electrolysis of potassium acetate⁴³ and ethyl malonate⁷ in an ethylene-glycol solution have also shown that only slight variations occurred in the yield of the principal product in the presence of foreign cations and anions, and when the anodic current densities, concentration, temperature, and anode material were varied.

TABLE 4.

| Cation | Na ⁺ | K ⁺ | NH ₄ ⁺ | Ca ²⁺ | Sr ²⁺ | Ba ²⁺ | Mg ²⁺ | Zn ²⁺ | Ni ²⁺ | UO ₂ ²⁺ |
|--------------------|-----------------|----------------|------------------------------|------------------|------------------|------------------|------------------|------------------|------------------|-------------------------------|
| Yield of ethane, % | 80 | 80 | 86 | 85 | 82 | 75 | 70 | 71 | 74 | 78 |

* Some acetamide, methylamine, and urea¹⁵⁴ were also found on electrolysis of ammonium acetate solution.

** In some cases, a theoretical yield of ethane was obtained on electrolysis of zinc acetate solution⁶³.

TABLE 5.

| pH | Yield of ethane, % | pH | Yield of Diethyl succinate, % |
|-----|--------------------|-----|-------------------------------|
| 2.3 | 67 | 1.4 | 70 |
| 4.3 | 70 | 2.9 | 75 |
| 4.7 | 77 | 3.5 | 74 |
| 5.2 | 74 | 4.8 | 74 |
| 5.5 | 71 | 7.0 | 70 |
| 11 | 50 | 8.3 | 64 |
| — | — | 10 | 59 |

TABLE 6.

| Acetate ions, M | Yield of ethane, % | Ethyl-malonate ions, M | Yield of diethyl succinate, % |
|-----------------|--------------------|------------------------|-------------------------------|
| 0.2 | 44 | 0.2 | 30 |
| 1.0 | 64 | 0.5 | 68 |
| 2.0 | 77 | 1.0 | 70 |
| 4.0 | 77 | 2.0 | 74 |
| | | 4.0 | 78 |

In large-scale electrolysis these advantages are offset by difficulties due to the low electrical conductivity of the solutions and high voltage required when working in non-aqueous solvents.

EFFECT OF THE STRUCTURE OF ACIDS ON THE COURSE OF THE KOLBE SYNTHESIS

Numerous investigations have shown that the yield of the product in the Kolbe reaction is not only affected by the conditions of electrolysis mentioned above but also to a considerable degree by the structure of the acids condensed. Thus, α -substituted carboxylic acids, cannot, in general, be subjected to the Kolbe reaction or enter into the reaction only with great difficulty. The same applies to α , β - and β , γ -unsaturated acids. Consequently, the structure of the initial acids or carboxylic-acid esters has to be considered before using them in electrolytic-condensation reactions.

The behaviour of various acids, in relation to their structure, during electrosynthesis reactions is described below. The data (unless otherwise indicated) refer to the electrolysis of aqueous solutions of sodium or potassium salts of carboxylic acids, usually in the presence of free acids.

1) Saturated Straight-Chain Acids

Fatty acids with straight C-chains of varying lengths have been subjected to the Kolbe electrolytic-condensation reaction, and normal synthesis products in satisfactory yields have been obtained in all cases, with the exception of formic acid. Many investigators have failed in attempts to produce hydrogen as well as CO₂ at the anode, according to the Kolbe reaction, on electrolysis of formic acid and its salts. CO₂ or CO and oxygen usually appear at the anode and the hydrogen formed is oxidised^{19,64-76} to water.

Some metals of the platinum group (rhodium⁷⁷⁻⁷⁹, palladium⁸⁰⁻⁸¹ and especially osmium⁸²) cause formic acid to decompose as follows:



Although this reaction suggests that electrodes made of these metals could be used in the electrolysis of formic acid, in actual experiments hydrogen did not appear at the anodes⁸³.

It has been mentioned above that yields of ethane (over 10%)* are obtained on electrolyzing acetic acid or its salts under similar conditions. The values of the anode potentials have been measured in numerous investigations^{57,41,85,86}. Experiments have been carried out on the electrolysis of acetates with alternating current^{76,78,87-91}, on the electrolysis of molten acetates^{88,92-95}, and on the acetates of various metals^{81,96-102}. The electrolytic condensation of deuterated acetates and acetic acid, in ordinary and heavy water¹⁰³⁻¹⁰⁸, has also been studied.

Many investigations have been devoted to the electrolysis of the following fatty acids: propionic^{10,16,19,35,40,64,73,86,87,94,95,106-109}, butyric^{10,11,20,26,35,64,87,95,110}, valeric^{1,24}, caproic^{24,62,87,111-114}, oenanthic^{14,111,112}, and caprylic^{112,115} **.

On electrolysis of potassium laurate in aqueous alcohol²³ *** n-docosane $C_{22}H_{46}$ was obtained in 90% yield; under analogous conditions myristic and palmitic acids^{28,117} gave the corresponding n-hexacosane (up to 89% yield) and n-triacontane (up to 88% yield). Pentadecanol and pentadecane are formed as by-products in the electrolysis of palmitic acid^{21,22}. Stearic acid gives a 73.6% yield of n-tetra-triacontane²⁸; other workers^{118,119} have also investigated this reaction.

2) Half-Esters of Dibasic Acids

Earlier it was shown that alkali-metal salts of dicarboxylic-acid monoesters react, on electrolysis, similarly to monobasic fatty acids and give diesters in high yields; an exception is the monoester of oxalic acid¹²⁰ which yields hydrogen, carbon dioxide, and ethylene on electrolysis.

The method proved to be very satisfactory for the production of up to C_{24} normal dicarboxylic acids. The optimum conditions are similar to those for the synthesis of ethane: high concentration of the electrolyte solution, low temperature, high anodic current density, etc. Table 7 shows the yields obtained.

3) Branched-Chain Acids and Half-Esters

The position of the substituent in alkyl-substituted acids affects the yield of the end product of the Kolbe reaction to a considerable extent. Branching of the chains in the α -position to the carboxylic group gives rise to the formation of olefins, alcohols, and other by-products; on electrolysis such acids or half-esters give a small yield or no yield of the required product. Thus, isobutyric acid gives only small quantities of 2,3-dimethylbutane, the predominant product being propylene; isopropyl alcohol and isopropyl butyrate are formed at the same time^{28,36,64}.

* In some cases⁸⁴ the electrolysis of an acetic acid solution differed from the usual course of the process.

** The yield of the principal condensation product has not been given.

*** In one investigation methyl laurate was obtained on electrolysis of potassium laurate in boiling methanol, in the presence of potassium chloride (graphite anode, 3.5 A, 90 V)¹¹⁶.

Electrolysis of α -methyl butyric acid gives negligible quantities of 3,4-dimethylhexane and leads to formation of but-2-ene, trimethyl carbinol, and its α -methylbutyric-acid ester²⁸ *.

Pivalic acid, in an analogous manner, yields predominantly isobutylene and but-2-ene and only insignificant quantities of hexamethylethane⁸⁴.

Swann¹⁵² was unable to identify the desired reaction product on electrolyzing α -methylvaleric acid and α -ethylbutyric acid. Later, more accurate investigations¹⁵³ on the electrolytic condensation of the latter acid made it possible to separate tetraethylethane together with other compounds (pent-2-ene, pentan-2-ol, pentan-3-one, etc.). It has also been shown that 17,18-dibutyltetra-triacontane can be prepared by the electrolysis of α -butylstearic acid¹⁵⁴.

Monoesters of dicarboxylic acids, containing alkyl substituents in the α -position to the carboxyl group, behave similarly. Thus, the ethyl esters of methylmalonic, ethylmalonic, dimethylmalonic¹²⁰, diethylmalonic^{120,155}, and dipropylmalonic¹⁵⁶ acids give on electrolysis small yields of the products.

It is possible to prepare the diethyl ester of β, β, β' -tetramethyl adipic acid (5-8% yield) by electrolyzing the monoethyl ester of α, α -dimethyl succinic acid in methanol¹⁵⁷. More satisfactory results are obtained with β, γ, δ -, etc. mono- and dialkyl-substituted acids. Table 8 gives data on the Kolbe reaction with acids and esters containing a branched chain in positions other than the α -position.

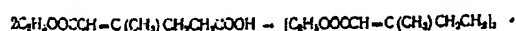
The optical activity of the asymmetric carbon atom in the β -position to the carboxyl group is preserved during the Kolbe reaction. It is therefore possible to use the electrochemical method to prepare stereoisomers of the dimethyl ester of β, β' -dimethylsuccinic acid, with an approximate yield of 70%, from the L(+) and D(-)-methyl esters of β -methylglutaric acid¹⁵⁸.

4) Unsaturated Acids

Numerous investigations have proved that neither α, β - nor β, γ -unsaturated acids undergo electrolytic condensation. The former acids give small quantities of acetylenes on electrolysis; thus acetylene is formed from acrylic acid¹⁶⁴, and allylene and but-2-yne respectively from crotonic^{165,166} and tiglic acids¹⁶⁷. Negative results were obtained on electrolysis of fumarates^{120,168}, maleates^{120,169}, citraconates¹²⁰, esters of allylmalonic¹²⁰ and muconic acids¹⁶⁹.

Fichter and Holbro¹⁷⁰ carried out experiments on the electrolytic condensation of β, γ -hexenoic acid and obtained negative results; they ascertained at the same time that γ, δ -hexenoic acid gave the normal condensation product deca-2,8-diene, in addition to by-products (penta-1,3-diene, pent-2-ene-6-ol, and its hexenol-acid ester).

The normal condensation product was obtained (55% yield) on electrolysis of 4-ethoxycarbonyl-3-methylbut-3-enecarboxylic acid¹⁶³ in methanol:



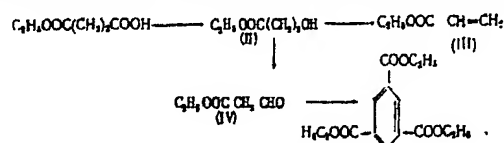
Satisfactory results were also obtained with unsaturated acids in which the double bond is further removed from the

* The inactive 3,4-dimethylhexane formed on electrolysis of d- and l- α -methylbutyric acids¹⁵⁰, together with the inactive butan-2-ol and its α -methylbutyric acid ester¹⁵¹.

TABLE 7. Kolbe reaction for half-esters of dicarboxylic acids with a straight $\text{ROOC}(\text{CH}_2)_n\text{COOH}$ chain.

| Initial half-ester | | | Solvent | Main product | By-products | Yield of main product, % | Reference |
|---|----|-----------------------|---|---|--|--------------------------|-----------|
| R | n | | | | | | |
| C ₂ H ₅ | 1 | Ethyl malonate | H ₂ O | H ₃ C ₇ OOC(CH ₂) ₂ COOC ₂ H ₅ | — | 60 | 5 |
| CH ₃ | 2 | Methyl succinate | CH ₃ OH | H ₃ C ₉ OOC(CH ₂) ₂ COOCH ₃ | ester C ₁₃ H ₂₂ O ₄ , probably CH ₃ COOCH ₃ H ₃ COOC(CH ₂) ₂ CH CH ₃ CH ₂ COOCH ₃ | 76 | 121 |
| C ₂ H ₅ * | 2 | Ethyl succinate | H ₂ O | H ₃ C ₉ OOC(CH ₂) ₂ COO ₂ C ₂ H ₅ | — | 35 | 6 |
| C ₆ H ₅ CH ₂ | 2 | Benzyl succinate | CH ₃ OH | C ₆ H ₅ CH ₂ OOC(CH ₂) ₂ COOCH ₂ C ₆ H ₅ | — | 50 | 122 |
| CH ₃ | 3 | Methyl glutarate | CH ₃ OH | H ₃ COOC(CH ₂) ₃ COOCH ₃ | — | 50 | 123 |
| C ₂ H ₅ | 3 | Ethyl glutarate | H ₂ O | H ₃ C ₉ OOC(CH ₂) ₃ COOC ₂ H ₅ | — | 62 | 125 |
| C ₆ H ₅ CH ₂ | 3 | Benzyl glutarate | CH ₃ OH + + C ₆ H ₅ N | C ₆ H ₅ CH ₂ OOC(CH ₂) ₃ COOCH ₂ C ₆ H ₅ | — | 28 | 5 |
| CH ₃ | 4 | Methyl adipate | CH ₃ OH | H ₃ COOC(CH ₂) ₄ COOCH ₃ | — | 24 | 126 |
| C ₂ H ₅ | 4 | Ethyl adipate | CH ₃ OH | H ₃ C ₉ OOC(CH ₂) ₄ COO ₂ C ₂ H ₅ | — | 51—61 | 127 |
| CH ₃ | 5 | Methyl pimelate | CH ₃ OH | H ₃ COOC(CH ₂) ₅ COOCH ₃ | — | 74.3 | 120 |
| C ₂ H ₅ | 5 | Ethyl pimelate | H ₂ O | H ₃ C ₉ OOC(CH ₂) ₅ COOC ₂ H ₅ | CH ₂ =CH(CH ₂) ₂ COOC ₂ H ₅ | 20 | 116** |
| CH ₃ | 6 | Methyl suberate | CH ₃ OH | H ₃ COOC(CH ₂) ₆ COOCH ₃ | — | — | 120 |
| C ₂ H ₅ | 6 | Ethyl suberate | H ₂ O | H ₃ C ₉ OOC(CH ₂) ₆ COOC ₂ H ₅ | — | 47 | 127 |
| CH ₃ | 7 | Methyl azelate | CH ₃ OH | H ₃ COOC(CH ₂) ₇ COOCH ₃ | H ₃ C ₉ OOC(CH ₂) ₇ CH=CH ₂ | — | 130 |
| C ₂ H ₅ | 7 | Ethyl azelate | H ₂ O | H ₃ C ₉ OOC(CH ₂) ₇ COOC ₂ H ₅ | — | — | 131 |
| CH ₃ | 8 | Methyl sebacate | CH ₃ OH | H ₃ COOC(CH ₂) ₈ COOCH ₃ | — | 44 | 127 |
| C ₂ H ₅ | 8 | Ethyl sebacate | C ₂ H ₅ OH + H ₂ O | H ₃ C ₉ OOC(CH ₂) ₈ COOC ₂ H ₅ | CH ₂ =CH(CH ₂) ₆ COOC ₂ H ₅ | — | 132 |
| CH ₃ | 9 | Methyl undecanoate | CH ₃ OH | H ₃ COOC(CH ₂) ₉ COOCH ₃ | — | 25 | 5 |
| C ₂ H ₅ | 9 | Ethyl undecanoate | H ₂ O | H ₃ C ₉ OOC(CH ₂) ₉ COOC ₂ H ₅ | — | 54 | 127 |
| CH ₃ | 10 | Methyl dodecanoate | CH ₃ OH | H ₃ COOC(CH ₂) ₁₀ COOCH ₃ | — | 21 | 133 |
| C ₂ H ₅ | 10 | Ethyl dodecanoate | H ₂ O | H ₃ C ₉ OOC(CH ₂) ₁₀ COOC ₂ H ₅ | CH ₂ =CH(CH ₂) ₈ COOC ₂ H ₅ | — | 134 |
| C ₂ H ₅ | 11 | Ethyl tridecanoate | H ₂ O | H ₃ C ₉ OOC(CH ₂) ₁₁ COOC ₂ H ₅ | — | 38—40 | 135 |
| CH ₃ | 12 | Methyl tetradecanoate | CH ₃ OH | H ₃ COOC(CH ₂) ₁₂ COOCH ₃ | — | 45 | 136 |
| C ₂ H ₅ | 12 | Ethyl tetradecanoate | C ₂ H ₅ OH + + H ₂ O | H ₃ C ₉ OOC(CH ₂) ₁₂ COOC ₂ H ₅ | CH ₂ =CH(CH ₂) ₁₀ COOC ₂ H ₅ | — | 132 |
| CH ₃ | 13 | Methyl pentadecanoate | CH ₃ OH | H ₃ COOC(CH ₂) ₁₃ COOCH ₃ | — | 42 | 127 |
| C ₂ H ₅ | 13 | Ethyl pentadecanoate | CH ₃ OH | H ₃ COOC(CH ₂) ₁₃ COOC ₂ H ₅ | — | 08 | 137, 138 |
| CH ₃ | 14 | Methyl hexadecanoate | H ₂ O | H ₃ C ₉ OOC(CH ₂) ₁₄ COOCH ₃ | — | 20 | 5 |
| C ₂ H ₅ | 14 | Ethyl hexadecanoate | H ₂ O | H ₃ C ₉ OOC(CH ₂) ₁₄ COOC ₂ H ₅ | CH ₂ =CH(CH ₂) ₁₂ COOC ₂ H ₅ | 20—40 | 120 |
| CH ₃ | 15 | Methyl heptadecanoate | H ₂ O | H ₃ COOC(CH ₂) ₁₅ COOCH ₃ | — | 30—35 | 138 |
| C ₂ H ₅ | 15 | Ethyl heptadecanoate | H ₂ O | H ₃ C ₉ OOC(CH ₂) ₁₅ COOC ₂ H ₅ | CH ₂ =CH(CH ₂) ₁₃ COOC ₂ H ₅ H ₃ C ₉ OOC(CH ₂) ₁₃ COOC ₂ H ₅ | 18—25 | 140 |
| CH ₃ | 16 | Methyl octadecanoate | H ₂ O | H ₃ COOC(CH ₂) ₁₆ COOCH ₃ | — | — | 141 |
| C ₂ H ₅ | 16 | Ethyl octadecanoate | H ₂ O | H ₃ C ₉ OOC(CH ₂) ₁₆ COOC ₂ H ₅ | — | 44—48 | 142 |
| CH ₃ | 17 | Methyl nonadecanoate | H ₂ O | H ₃ COOC(CH ₂) ₁₇ COOCH ₃ | — | 16.5 | 143 |
| C ₂ H ₅ | 17 | Ethyl nonadecanoate | H ₂ O | H ₃ C ₉ OOC(CH ₂) ₁₇ COOC ₂ H ₅ | — | — | 144 |
| CH ₃ | 18 | Methyl eicosanoate | CH ₃ OH | H ₃ COOC(CH ₂) ₁₈ COOCH ₃ | — | 32 | 145 |
| C ₂ H ₅ | 18 | Ethyl eicosanoate | CH ₃ OH | H ₃ C ₉ OOC(CH ₂) ₁₈ COOC ₂ H ₅ | — | — | 146 |
| CH ₃ | 19 | Methyl heneicosanoate | CH ₃ OH | H ₃ COOC(CH ₂) ₁₉ COOCH ₃ | — | 65 | 137 |
| C ₂ H ₅ | 19 | Ethyl heneicosanoate | CH ₃ OH | H ₃ C ₉ OOC(CH ₂) ₁₉ COOC ₂ H ₅ | — | — | 146 |
| CH ₃ | 20 | Methyl docosanoate | CH ₃ OH | H ₃ COOC(CH ₂) ₂₀ COOCH ₃ | CH ₂ =CH(CH ₂) ₁₈ COOCH ₃ | 43 | 147 |
| C ₂ H ₅ | 20 | Ethyl docosanoate | CH ₃ OH | H ₃ C ₉ OOC(CH ₂) ₂₀ COOC ₂ H ₅ | CH ₂ =CH(CH ₂) ₁₈ COOCH ₃ probably CH ₃ COOC(CH ₂) ₁₈ COOCH ₃ | — | 148 |
| CH ₃ | 21 | Methyl tricosanoate | H ₂ O + + C ₂ H ₅ OH | H ₃ C ₉ OOC(CH ₂) ₂₁ COOCH ₃ | CH ₂ =CH(CH ₂) ₁₉ COOC ₂ H ₅ | 41 | 139 |
| CH ₃ | 22 | Methyl tetracosanoate | CH ₃ OH | H ₃ COOC(CH ₂) ₂₂ COOCH ₃ | CH ₂ =CH(CH ₂) ₂₀ COOCH ₃ | — | 148 |
| C ₂ H ₅ | 22 | Ethyl tetracosanoate | H ₂ O + + C ₂ H ₅ OH | H ₃ C ₉ OOC(CH ₂) ₂₂ COOC ₂ H ₅ | — | 5.1 | 132 |
| CH ₃ | 23 | Methyl pentacosanoate | CH ₃ OH | H ₃ COOC(CH ₂) ₂₃ COOCH ₃ | — | — | 132 |
| C ₂ H ₅ | 23 | Ethyl pentacosanoate | CH ₃ OH | H ₃ C ₉ OOC(CH ₂) ₂₃ COOC ₂ H ₅ | H ₃ COOC(CH ₂) ₁₉ COOCH ₃ | 8 | 148 |
| CH ₃ | 24 | Methyl hexacosanoate | CH ₃ OH | H ₃ COOC(CH ₂) ₂₄ COOCH ₃ | — | 10 | 149 |
| C ₂ H ₅ | 24 | Ethyl hexacosanoate | CH ₃ OH | H ₃ C ₉ OOC(CH ₂) ₂₄ COOC ₂ H ₅ | — | — | 149 |
| CH ₃ | 25 | Methyl heptacosanoate | C ₂ H ₅ OH + + H ₂ O | H ₃ C ₉ OOC(CH ₂) ₂₅ COOCH ₃ | — | 4.7 | 132 |
| C ₂ H ₅ | 25 | Ethyl heptacosanoate | C ₂ H ₅ OH + + H ₂ O | H ₃ C ₉ OOC(CH ₂) ₂₅ COOC ₂ H ₅ | — | 01 | 149 |

* The cyclic ester of trimesic acid is formed in the electrolysis of monoethyl succinate¹². The authors assume that, the ester of β -hydroxypropionic (II) and acrylic (III) acids are formed according to the Hofer-Moest reaction; (II) is oxidised at the anode to the formalacetate (IV), which is then condensed to the trimesate (V):



** The electrolysis is carried out in the presence of sodium formate.

carboxyl group, such as undec-10-enoic^{112,165,171,172}, petroselinic¹⁷³, oleic^{165,173-175}, erucic¹⁷³, and ricinoleic^{175,176}.

The use of the Kolbe synthesis was further extended by Bounds et al.¹⁷¹, who found that the geometrical configuration at the double bonds was preserved during the electrolysis of acids of the ethylene series; thus unsaturated compounds with known geometrical configuration can be synthesised by the electrochemical method. In the electrolytic condensation of oleic and elaidic acids in absolute methanol the authors obtained the corresponding *cis-cis*-tetraatriaconta-9,25-diene (23% yield) and *trans-trans*-tetraatriaconta-9,25-diene (44% yield).

In the electrolytic condensation of unsaturated acids the anode generally becomes coated with a layer of insoluble polymers which are formed during the reaction and inhibit it, especially when the acids contain several double bonds¹⁷⁷. A large quantity of polymeric products separates during the electrolytic condensation of the monomethyl ester of dec-5-ene-1,10-dicarboxylic acid¹⁷⁸; only small yields of the normal product (eicosa-5,15-diene-1,20-dicarboxylic acid) were obtained.

5) Aliphatic Hydroxy- and Alkoxyacids

Aliphatic hydroxyacids are not very suitable for the Kolbe reaction as the hydroxyl group is readily oxidised, especially when it is adjacent to the carboxyl group. Alkoxyacids are more stable.

On electrolysis, α -hydroxy- and α -alkoxyacids (e.g. glycolic¹⁷⁹, methoxy- and ethoxyacetic acids¹⁸⁰⁻¹⁸², α -hydroxybenzoic acid¹⁸², lactic^{180,183,184}, α -hydroxybutyric, and α -hydroxyisobutyric acids¹⁸⁰) give an aldehyde or ketone as well as an acid containing one carbon atom less than the starting material. Lactic acid has thus been converted to pyruvic acid¹⁸⁵.

When the monoethyl esters of ethoxymalonic and symmetrical diethoxysuccinic acid were electrolysed¹⁸², only small quantities of the condensation product (the diethyl ester of diethoxysuccinic acid) were separated in the first case, and none at all in the second case.

Negative results were obtained on electrolysis of β -hydroxyacids; electrolytic condensation of hydracrylic acid^{180,186}, for example, only resulted in the formation of oxidation products (formic acid, CO₂, etc.). Formic acid, crotonaldehyde, and other oxidation products are obtained from β -hydroxybutyric acid¹⁸⁰.

β -Alkoxyacids undergo the Kolbe reaction; the diisopentyl ether of butane-1,4-diol (50% yield) was prepared by electrolysis of β -isopentyloxypropionic acid^{187,188*}.

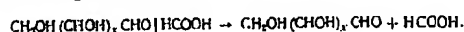
In the electrolysis of γ -pentyloxybutyric acid¹⁸⁰ it was found that hydrolysis replaced the Kolbe reaction, leading to the formation of pentyl alcohol and γ -hydroxybutyric acid. In the electrolysis of γ -ethoxybutyric acid¹⁹¹ it was possible to separate the diethyl ether of hexanediol (20% yield). Further electrolytic condensation experiments have been carried out with 4,8-dimethyl-9-ethoxynonanolic acid in anhydrous methanol¹⁵⁸ and 2,6,11,15-tetramethyl-1,16-diethoxyhexadecane was formed.

* Experiments on a similar reaction via the thermal decomposition of di(β -isopentyloxypropionyl) peroxide were carried out by Fichter and Schnider¹⁸⁹.

Satisfactory results were obtained by electrolysis of water-alcohol and alcoholic solutions of ricinoleic^{175,176} and dihydroxystearic acids¹⁷⁰. The electrolysis of (*±*)-*threo*-9,10-dihydroxystearic and (*±*)-*erythro*-9,10-dihydroxystearic acid in absolute alcohol¹⁷¹ gave the corresponding diastereoisomers of the tetrol in 50% and 20% yields respectively.

Attention should also be drawn to electrolytic condensation experiments with glyceric¹⁸⁰, β -methylglyceric, and β -methylglycidic acids¹⁸³, as well as with the *d*-tartaric-acid monoethyl ester¹⁸³; they all gave negative results, as only oxidation products of the initial acids were formed.

In 1908 Neuberg¹⁸⁴ attempted to produce polyglycols by the electrolysis of aldonic acids (*d*-gluconic, *l*-arabonic, *i*-erythronic, and *i*-glyceric acid) and glycolic acid; however, aldehydes, apparently the decomposition products of the initial acids, were separated instead of the required condensation products:



By this method glucose can be converted stepwise to formaldehyde. It can be assumed that keto-acids of the type $\text{CH}_2\text{OH}(\text{CHOH})_x\text{COCOOH}$ are formed as by-products.

Electrolytic decomposition has also been used in conjunction with other carboxylic-acid derivatives of sugars, e.g. melibionc acid¹⁸⁵.

6) Aldo-Acids

Free aldo-acids are readily oxidised in the presence of oxygen and therefore, cannot be subjected to electrolytic condensation. However, their derivatives, the acetals, are stable under electrolysis conditions.

On electrolysis of an aqueous solution of the potassium salt of β,β -diethoxypropionic acid (Pt anode, Ni cathode, 3A, 8-10V) a 60% yield of bis-diethylacetal of succinaldehyde^{186*} was obtained. Analogous experiments with the potassium salt of γ,γ -diethoxybutyric acid led to the formation of the bis-diethylacetal of adipaldehyde (35% yield); acrolein acetal was also formed in this reaction.

The bis-dimethylacetal of hexadecane-1,16-dial was obtained by electrolysis of 9,9-dimethoxynonanolic acid in anhydrous methanol (3-4 A, 40-50 V, 50°-55°); under similar conditions 10,10-dimethoxydecanoic acid yielded the bis-dimethylacetal of octadecane-1,18-dial¹⁸⁸.

It is very difficult to prepare dialdehydes by purely chemical methods.

7) Keto-Acids

In the electrolytic condensation of keto-acids, the yield of the end product depends to a considerable degree on the position of the CO group in the molecule of the initial acid. Thus pyruvic acid, the simplest keto-acid, gives only small yields of butane-2,3-dione; the principal reaction consists in the oxidation of the keto-acid to acetic acid and carbon dioxide¹⁸⁹. Electrolysis of a water-alcohol solution of pyruvic acid, in an electrolyser comprising a diaphragm,

* The electrolysis of β,β -diethoxypropionic acid was described by Japanese investigators¹⁹¹.

TABLE 8. Kolbe reaction for acids and half-esters with branched chains.

| Initial reactant | Solvent | Main product | By-products | Yield of main product | Reference |
|--|---|--|---|-----------------------|-----------|
| Isovaleric acid | H ₂ O | $\begin{array}{c} \text{H}_3\text{C} \\ \\ \text{H}_3\text{C}-\text{CH}-\text{CH}_2-\text{CH}_2-\text{CH}_2-\text{CH}_3 \\ \\ \text{CH}_3 \end{array}$ | Isobutyl isovalerate, trimethyl carbinol, isobutyl alcohol, isobutyl aldehyde, 2-methylpropane, but-2-ene | — | 24 |
| Isocaproic acid | " | $\begin{array}{c} \text{H}_3\text{C} \\ \\ \text{H}_3\text{C}-\text{CH}-(\text{CH}_2)_4-\text{CH}-\text{CH}_3 \\ \\ \text{CH}_3 \end{array}$ | — | 20 | 153 |
| 4,8-Dimethyl-9-ethoxynonanolic acid | CH ₃ OH | $[\text{C}_6\text{H}_5\text{OCH}_2\text{CH}(\text{CH}_3)(\text{CH}_2)_3\text{CH}(\text{CH}_3)\text{CH}_2\text{CH}_3]$ | — | — | 158 |
| 10-Methyloctadecanoic acid | H ₂ O + C ₂ H ₅ OH | $\text{CH}_3(\text{CH}_2)_7\text{CH}(\text{CH}_3)(\text{CH}_2)_6\text{CH}(\text{CH}_3)(\text{CH}_2)_7\text{CH}_3$ | — | 8,3 | 110 |
| Ethyl 3,3-dimethyl succinate * | CH ₃ OH | $\text{H}_3\text{C}_2\text{OOC}-\text{C}(\text{CH}_3)_2\text{CH}_2\text{CH}_2\text{C}(\text{CH}_3)_2\text{COOC}_2\text{H}_5$ | Ethyl β , β -dimethylacrylate, diethyl ester of asymmetric dimethylsuccinic acid, monoethyl ester of asymmetric dimethylsuccinic acid | 32—35 | 157 |
| Methyl β -methylglutarate | CH ₃ OH + H ₂ O | $\text{CH}_3\text{OOCCH}_2\text{CH}(\text{CH}_3)\text{CH}_2\text{CH}_2\text{CH}(\text{CH}_3)\text{CH}_2\text{COOCH}_3$ | — | 80 | 159 |
| Ethyl β , β -dimethylglutarate | H ₂ O | $\text{H}_3\text{C}_2\text{OOCCH}_2\text{CH}(\text{CH}_3)\text{CH}_2\text{CH}_2\text{C}(\text{CH}_3)_2\text{CH}_2\text{COOC}_2\text{H}_5$ | — | 42 | 160 |
| Methyl β , β -dimethylglutarate | CH ₃ OH + pet. ether | $\text{H}_3\text{COOCCH}_2\text{CH}(\text{CH}_3)\text{CH}_2\text{CH}_2\text{C}(\text{CH}_3)_2\text{CH}_2\text{COOCH}_3$ | — | 75 | 161 |
| Methyl methyladipate | CH ₃ OH | Dimethyl dimethyladipate | — | 78 | 162 |
| Ethyl γ -methyladipate | CH ₃ OH | $\text{H}_3\text{C}_2\text{OOCCH}_2\text{CH}(\text{CH}_3)(\text{CH}_2)_3\text{CH}(\text{CH}_3)\text{CH}_2\text{COOC}_2\text{H}_5$ | — | 70 | 163 |
| Methyl β , β , β' , β' -tetramethylsuccinate | CH ₃ OH + pet. ether | $[\text{H}_3\text{COOC}-\text{CH}_2-\text{C}(\text{CH}_3)_2(\text{CH}_2)_2\text{C}(\text{CH}_3)_2\text{CH}_2]_2$ | — | 75 | 161 |
| Ethyl 3,7-dimethylnonano-1,9-dicarboxylate | CH ₃ OH | $[\text{H}_3\text{C}_2\text{OOCCH}_2\text{CH}_2\text{CH}(\text{CH}_3)(\text{CH}_2)_3\text{CH}(\text{CH}_3)\text{CH}_2\text{CH}_2]_2$ | Ethyl dimethyl-4,8-decanedioate, ethyl 4,8-dimethyldec-9-enoate | 48 | 164 |

* Here and subsequently the position of the substituent is given in relation to the free carboxyl group.

and graphite, platinum, or lead anodes, gave in all cases lactic²⁰⁰ as well as acetic acid.

The monoethyl ester of acetone dicarboxylic acid²⁰¹ has been electrolysed yielding the diester of hexane-2,5-dione-1,6-dicarboxylic acid (12% yield).

Considerably better results were obtained in the electrolysis of aqueous solutions of levulinic acid¹⁹⁹ in which the normal product, octane-2,7-dione, was obtained in 50% yield. When this process is carried out in a 25% sulphuric-acid solution or in methanol, the yield increases to 66% and 70% respectively²⁰².

The electrosynthesis process is also favoured by increasing the distance between the keto- and the carboxyl group in the keto-acid molecule. Thus tetradecane-3,12-dione, the normal product, was obtained in 74.8% yield on electrolysis of δ -propionylvaleric acid²⁰³.

8) Halogeno-Substituted Acids

Up to 1956 very little was published on the possibility of using the Kolbe reaction for the synthesis of halogeno-substituted compounds. Kaufler and Herzog²⁰⁴ discovered in 1909 that electrolysis of chloro-, bromo- and iodoacetic acids did not give products of the Kolbe reaction. Negative results were also obtained in the electrolysis of dichloro-^{204,205}, difluorochloro-²⁰⁷, and trichloroacetic²⁰⁸⁻²¹¹ acids. Only the electrolysis of trifluoroacetic acid gave

traces of hexafluoroethane^{207,212}. Unsatisfactory results were also obtained in attempts at synthesis from α , α -dichloropropionic²⁰⁸, β -iodopropionic²⁰⁸, 2,2,3-trichlorobutyric²¹³, and d,l - α -bromopropionic acids²¹⁴. Japanese investigators¹⁷⁷ were also unsuccessful when they electrolysed tetrabromostearic acid in ethanol.

In 1953 Korsching²¹⁵ succeeded in isolating the normal product (1,20-dibromoeicosane, 20% yield) in the electrolysis of 11-bromoundecanoic acid; later²¹⁶ he succeeded in preparing 1,10-dibromodecane from ω -bromocaproic acid in a water-alcohol solution.

Detailed investigations on the electrolytic condensation of ω -halogenocarboxylic acids, carried out in 1956,¹⁵⁷ showed that normal products are obtained when β -chloropropionic, δ -chlorovaleric, 6-chlorohexanoic and 10-chlorodecanoic are electrolysed in absolute methanol (11–16 A/dm², 2 A, 110 V, 50°), in contrast to the reaction for chloroacetic and γ -chlorobutyric acid.

Halogen is split off during the electrolysis of lower ω -bromocarboxylic acids. Normal condensation reactions could only be observed with acids containing 11 or more carbon atoms in the chain. All ω -iodocarboxylic acids (containing 2–11 carbon atoms in the molecule) decompose on electrolysis. ω -Fluorocarboxylic acids give considerably better results. The lower members of this series (fluoroacetic, β -fluoropropionic, and γ -fluorobutyric acid) do not undergo the Kolbe reaction; the glass electrolyser is attacked – an indication that hydrogen fluoride is evolved in the process. However, δ -fluorovaleric, 6-fluorohexanoic, 7-fluoroheptanoic, 8-fluorooctanoic, 9-fluorononanoic, 10-fluorodecanoic, and 11-fluoroundecanoic acids give satisfactory yields of ω , ω' -difluoroalkanes.

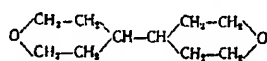
* See also refs. 92, 205, and 206 on the electrolysis of chloroacetic acid.

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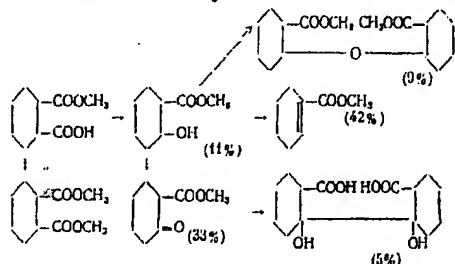
On electrolysis of the monomethyl ester of *cis*-cyclopropane-1,2-dicarboxylic acid²³² ring cleavage occurs and the ester of hexa-2,4-diene-1,6-dicarboxylic acid* is formed. Negative results were obtained on electrolysis of cyclobutanecarboxylic acid^{231,234}, cyclohexanecarboxylic acid^{231,235}, and *p*-methylcyclohexanecarboxylic acid²³⁶.

Later investigations of the electrolysis of cyclohexanecarboxylic acid showed that negligible quantities of the anodic-synthesis product, bi(cyclohexyl), are formed when the process is carried out in an acidic solution²³⁷; when a methanol-pyridine mixture is used as solvent, the yield of bi(cyclohexyl) is 14%.²³⁸

Satisfactory yields (20%) of 4,4-bis-tetrahydropyran were obtained by electrolysis of an aqueous solution of the potassium salt of tetrahydropyran-4-carboxylic acid²³⁹; at the same time equal quantities of 4-hydroxytetrahydropyran, and its tetrahydropyran-4-carboxylic-acid ester were isolated.

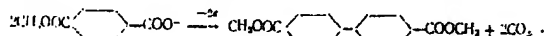


An aqueous solution of the monoethyl ester of cyclohexylmalonic acid gave, on electrolysis, α, α' -dicyclohexylsuccinic acid (a mixture of the two stereoisomers); the yield was 17%.¹²³ Detailed studies of the electrolysis of hexahydrophthalate showed that a number of by-products were formed (see below); the expected anodic-synthesis product was not obtained either in aqueous or in methanol solutions²⁴⁰



Later²⁴¹ it was proved that under the given conditions of electrolysis of monomethyl esters of *cis*- and *trans*-hexahydrophthalic acid in anhydrous methanol, normal condensation products, i.e. dimethyl *trans-anti-trans*- and *trans-syn-trans*-bi(cyclohexyl)-2,2'-dicarboxylate are formed, but only in small quantities; the principal reaction products are monomeric compounds - the methyl esters of cyclohex-1-ene-1-carboxylic, cyclohex-2-ene-1-carboxylic acid, and cyclohexanecarboxylic acids.

It should be pointed out that the isomeric methyl ester of *trans*-1,4-cyclohexanedicarboxylic acid gave appreciable quantities (32%) of the Kolbe-synthesis product²⁴²



Ester groups in the *p*-position obviously aid the electrolytic condensation.

Electrolytic condensation of several alicyclic dicarboxylic-acid esters was successfully carried out in a mixture of methanol and petroleum ether (Pt anode, 4-6 A/dm²). Thus, electrolysis of the monomethyl esters of cyclopentane-1,1-diacetic acid, 2,2,5,5-di(tetramethylene)hexane-1,6-dicarboxylic acid, cyclohexane-1,1-diacetic acid, and

* In the electrolysis of the monomethyl ester of cyclopropane-1,2-dicarboxylic acid in methanol the ester of the unsaturated dibasic acid, probably α, α' -dimethylenedipic acid, was formed instead of the expected diester of dicyclopropanedicarboxylic acid¹²³.

2,2,5,5-di(pentamethylene)hexane-1,6-dicarboxylic acid¹⁶¹ gave the dimethyl esters of the corresponding 2,2,5,5-di(tetramethylene)hexane-1,6-dicarboxylic acid, 2,2,5,5,8,8,11,11-tetra(tetramethylene)dodecane-1,12-dicarboxylic acid, 2,2,5,5-di(pentamethylene)hexane-1,6-dicarboxylic acid (38% yield), and 2,2,5,5,8,8,11,11-tetra(pentamethylene)dodecane-1,12-dicarboxylic acid (14.5% yield).

Other investigations dealt with the electrolytic condensation of half-esters of camphoric acid²⁴³⁻²⁴⁹.

12) Aromatic and Aliphatic-Aromatic Acids

For a long time it was assumed that salts of aromatic acids, unlike those of aliphatic acids, did not enter into the Kolbe reaction.

Experiments on the electrolytic condensation of solutions of salts of benzoic acid in water^{92,250,251}, sulphuric acid²⁵², and absolute methanol²⁵³, and the electrolysis of fused benzoic acid^{23,25} were indeed unsuccessful. The reaction products were, in general, benzene, free benzoic acid, oxygen, etc. The same was true of the electrolytic condensation of substituted benzoic acids such as *o*-nitrobenzoic²⁵⁴, *p*-nitrobenzoic²¹⁷⁻²⁵⁵, *p*-toluic^{256,258,257}, opianic²⁵⁸, salicylic, *m*-hydroxybenzoic, and *p*-hydroxybenzoic acids⁹², 3,5-dimethyl- and 2,4-dimethylbenzoic acids²⁵⁹, and ethyl phthalate¹²⁰.

On electrolysis 2,5-dimethyl-4-*t*-butylbenzoic acid in methanol²⁵³ one of the methyl substituents in the aromatic nucleus was oxidised and, on subsequent esterification, the methyl ester of 6-methyl-4-*t*-butyl-1,2-phthalic acid was formed together with 1,3-dimethyl-5-*t*-butylbenzene.

Negative results were obtained on electrolysis of some aromatic acids which contained the carboxyl group in the side chain, e.g. aqueous solutions of potassium salts of phenylacetic acid²⁶⁰, diphenylacetic acid²⁶¹, and also of the ethyl ester of benzylmalonic acid¹²⁰.

In the 1930's Fichter et al. refuted the conclusion that aromatic acids do not undergo the Kolbe reaction. They proved that the characteristic electrosynthesis product (2,2'-diindanyl) is formed in the electrolysis of indane-2-carboxylic acid in methanol^{262,264}, although only in small quantities. Other aromatic acids can be electrolysed successfully by using a methanol-pyridine mixture as solvent. Thus, benzoic, phenylacetic, β -phenylpropionic, and phenoxycetic acids²⁶⁵ yielded the corresponding diphenyl, dibenzyl (50% yield), 1,4-diphenylbutane (37% yield), and the diphenyl ester of ethylene glycol (45% yield). Normal products (the respective yields being 27, 24, and 8%) were obtained on electrolysis *o*-, *m*-, and *p*-cresoxyacetic acids²⁶⁷ under analogous conditions.

Under these conditions it was possible to prepare tetraphenylethane by electrolysis diphenylacetic acid^{268,269}, and also 1,4-diphenylbutane (34% yield), 1,6-diphenylhexane (37% yield), and 1,8-diphenyloctane (47% yield) from the corresponding β -phenylpropionic, γ -phenylbutyric, and δ -phenylvaleric acids²⁷⁰.

Dibenzyl was produced in 1952 by electrolysis phenylacetic acid in anhydrous methanol and also by electrolysis of fused phenylacetic acid²⁶⁶. At the same time it was found that electrolysis of phenylacetic acid in acetic acid

* The diethyl ester of dibenzylsuccinic acid¹⁶² is formed when the monoethyl ester of benzylmalonic acid is electrolysed in the presence of acetic acid.

and also of diphenyl- and triphenylacetic acids in acetic acid or anhydrous methanol caused the acetoxylation or methoxylation of the benzene nuclei respectively.

Results on the electrolysis of γ -phenylisovaleric, 3,3,3-triphenylpropionic, and 3,3,3-tri-*p*-t-butylphenylpropionic acid in methanol, using a platinum anode, were published recently²⁷¹. Electrolysis of the first acid, in the presence of an excess of sodium acetate (15 A/dm², 60°) gave *t*-pentylbenzene (32%); 2,5-diphenyl-2,5-dimethylhexane and homologues of α -methylstyrene were separated as by-products. Electrolysis of the second acid led to the formation of the phenyl ester of 3,3-diphenyl-3-methoxypropionic acid, and the third acid gave the *p*-*t*-butylphenyl ester of 3,3-di-*p*-*t*-butylphenyl-3-methoxypropionic acid.

13) Dicarboxylic Acids

Alkali-metal salts of dicarboxylic acids do not undergo the Kolbe reaction at the anode.

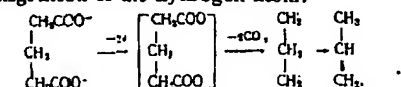
Only hydrogen, oxygen, carbon dioxide^{70,73-76,272-275}, and in some instances glycolic acid²⁷⁷ could be detected in the electrolysis products of oxalic acid.

The higher dibasic acids (malonic to sebacic) form olefins and alcohols which contain two carbon atoms less than the original acid, and also saturated and unsaturated acids with one carbon atom less. Attempts to isolate cycloalkanes, the formation of which might be expected as 2 molecules of carbon dioxide are split off simultaneously from the dibasic-acid anion, were unsuccessful.

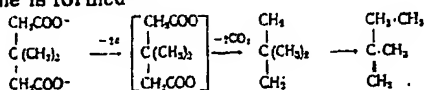
Ethylene, hydrogen, oxygen, carbon monoxide, and carbon dioxide are formed in the electrolysis of malonic acid^{272,278}. Schall²⁷⁹ found that in the electrolysis of the alkali-metal salts of malonic acid, hydrogen peroxide is formed on cooling, probably through the hydrolysis of the intermediate malonic-acid peroxide.

Electrolysis of succinic acid^{272,280} leads to the formation of ethylene, oxygen, hydrogen, and carbon dioxide*; tartaric and oxalic acid were also isolated as oxidation products²⁸². By electrolyzing succinic acid in an alkaline solution in the presence of sodium perchlorate, Hofer and Moest¹⁰ also prepared β -hydroxypropionic acid** together with acrylic acid, carbon monoxide, carbon dioxide, acetylene, acetaldehyde, methanol, and acetic and formic acids.

When an aqueous solution of the potassium salt of glutaric acid was electrolysed with the object of obtaining cyclopropane²⁸⁴, the products were found instead to be propylene, carbon dioxide and oxygen. The formation of propylene is due to the migration of the hydrogen atom:



In the electrolysis of the salt of β , β -dimethylglutaric acid, where no hydrogen transfer can take place, the whole hydrocarbon group is shifted and the unsymmetrical methyl-ethylethylene is formed¹⁸⁰



* See Fichter and Fritsch's publications²⁸¹ on the decomposition of succinic-acid peroxide to carbon dioxide and ethylene.

** β -hydroxypropionic acid and the acrylic acid formed from it were also obtained by the thermal decomposition of monopersuccinic acid²⁸³.

The cyclic product -- dimethylcyclopropane -- was not formed.

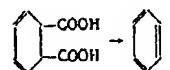
A number of higher dicarboxylic acids, such as adipic²⁸⁵, pimelic²⁸⁶⁻²⁸⁹, suberic²⁸⁹⁻²⁹⁰, and sebacic²⁷² acids, have been electrolysed but in no case could the formation of cycloalkanes be detected.

14) Substituted Dibasic Acids

On electrolysis, methylmalonic acid yields ethylene, hydrogen, oxygen, and carbon monoxide and dioxide^{92,272}; propylene, propanol, isopropyl alcohol, and other oxidation products²⁷⁸ are formed from ethylmalonic acid and pyrotartaric acid.

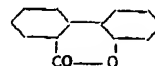
Salts of malic, tartaric, and pyrotartaric acids²⁹¹ have also been electrolysed; malic acid gave malonic acid, which was then readily oxidised to oxalic and formic acids and carbon dioxide. Formic acid was formed from the other two acids. Oxalic acid and carbon dioxide were obtained in the electrolysis of mesoxalic acid.

The preparation of benzene (70% yield) from *trans*-1,2-dihydrophthalic acid²⁸²,



and propadiene²³³ by opening the ring of 1,2-cyclopropane-dicarboxylic acid on electrolysis of its salts, are isolated instances of the formation of unsaturated hydrocarbons in the electrolysis of substituted dibasic acids.

In 1957²⁸³ *cis-syn-cis*-, *trans-anti-trans*-, and *trans-syn-trans*-perhydrodiphenic acids were electrolysed in anhydrous methanol. The principal reaction product (approx. 45%) was in all cases the corresponding isomeric γ -lactone; its structure was determined by synthesis:



15) Unsaturated Dibasic Acids

In the electrolysis of salts of fumaric²⁸⁰ and maleic^{180,280*} acids carbon dioxide and acetylene are formed at the anode, the maleic acid being partly converted to fumaric acid.

Acetylene, acrolein, carbon monoxide, and carbon dioxide are obtained when glutaric acid is electrolysed²⁸⁵.

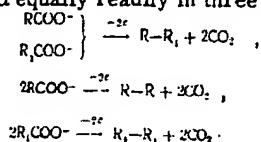
Mesaconic acid²⁸⁶ forms propyne, and acrylic and itaconic acids, whilst its isomers, itaconic²⁹⁷⁻²⁹⁹ and citraconic²⁸⁶ acids, yield propyne, and acrylic and mesaconic acids. The type of salt of the citraconic acid used in the electrolytic condensation determines the yield of propyne; thus the highest yield (52%) of propyne was obtained by electrolyzing the rubidium salt of citraconic acid and the lowest yield (16%) when the lithium salt was used³⁰⁰.

Fumaric acid and acetylene^{301,302} have been prepared from acetylenedicarboxylic acid.

* Maleic acid is formed in the electrolysis of hydroxymaleic acid²⁸⁴.

"CROSSED" ELECTROLYTIC CONDENSATION

In this reaction acids or half-esters of dicarboxylic acids with different radicals are caused to react. Such reactions do not take place very readily and are seldom carried out. The main difficulty lies in the fact that the condensation process can proceed equally readily in three principal ways:



As in the simple Kolbe synthesis a series of side reactions accompany the electrolysis; this explains the comparatively small yield of the main reaction product (R-R_1).

The reaction was first carried out by Würtz in 1855^{3,4}, who succeeded in preparing a small quantity of heptane and decane by electrolyzing a mixture of oenanthic and acetic acids and a mixture of oenanthic and valeric acids respectively*. It was later used by other investigators for synthesizing a number of compounds. Mixtures of acids and half-esters of dicarboxylic acids with straight and branched chains were used and also carboxylic acids containing various substituents (keto- and amino-groups, halogens, etc.) and double and triple bonds in the chain.

Most publications on the "crossed" electrolytic condensation recommend that methanol solutions of potassium or sodium salts of the acids be used as electrolytes**. Electrolysis of a mixture of the methyl ester of adipic acid with acetic acid or hexanoic acid in anhydrous methanol (11–16 A/dm², 1–2 A, 110 V, 50°) yielded the corresponding methyl esters of hexanoic acid (42% yield)*** and decanoic acid (61% yield)¹³⁷.

The methyl ester of arachidic acid (41.5% yield) was obtained from the monomethyl ester of adipic and palmitic acids in boiling methanol (7 A/dm², 5.6 A, 12–38 V)^{137,138}. Under analogous conditions the methyl ester of adipic and stearic acids yielded the methyl ester of behenic acid (38% yield); electrolysis in methanol (4–12 A/dm², 40°–60°) of the methyl ester of adipic acid with butyric, capric, or lauric acids yielded the corresponding caprylic (30.6% yield), myristic (7% yield), and palmitic acids (23.4% yield)³⁰⁴****.

The use of benzyl esters as starting materials in place of the ethyl and methyl esters in syntheses of this kind makes it possible to avoid esterification with the methanol used as solvent, which very often accompanies "crossed" syntheses. Thus, palmitic (30% yield) and arachidic (25% yield) acids respectively¹²⁴ can be produced by condensing the benzyl ester of succinic and myristic or stearic acids in methanol (6–7 A/dm², 1–1.3 A, $t < 50^\circ$) and hydrolysing the products.

According to other data¹²⁵ the use of monobenzyl esters as one component during the "crossed" electrolytic conden-

* Schorlemmer³⁰³ repeated Wurtz' experiment in 1865 and obtained similar results.

** Unless otherwise stated "crossed syntheses" discussed below refer to the electrolysis of sodium and potassium salts of carboxylic acids, carried out in the presence of free acids.

*** Hexanoic acid (40.8% yield) was also obtained in the electrolysis of the monoethyl ester of adipic acid and acetic acid³⁰⁴ in ethanol (10 A/dm², 0.8 A, 40°–60°).

**** See also publications by Japanese authors³⁰⁸ on the electrolysis of dibasic acid esters with fatty acids.

sation of a mixture of half-esters of dicarboxylic acids, facilitates the isolation of the product. Thus the methyl ester of pimelic acid (24% yield) was obtained by electrolyzing the benzyl ester of glutaric acid and the methyl ester of succinic acid in a methanol-pyridine mixture (Pt electrode, 0.2 A, 34 V, 35°) and subsequent hydrogenolysis of the reaction products. It has been shown³⁰⁴ that addition of tetrahydrofuran is advisable in this synthesis. Electrolysis of the ethyl ester of adipic acid with myristic, pentadecanoic, and stearic acids in methanol in the presence of tetrahydrofuran (2.5–12.5 A/dm², 0.2–1 A, 40°–60°) gave the corresponding stearic (12.7% yield), nonadecanoic (6.7% yield), and behenic acids (4.7% yield).

The same rule as in simple Kolbe reactions applies here for substituted and unsaturated acids: the starting compound must not contain a branched chain, a functional group, or an unsaturated bond in the α -position to the carboxyl group. It should be pointed out, however, that Japanese chemists³⁰⁷ succeeded in preparing 10-methyloctadecanoic acid by electrolyzing a mixture of the ethyl ester of sebacic and methyloctalacetic acids (reaction conditions and yields are not given).

3-methylundecanoic acid (46% yield) was obtained in the electrolysis of monomethyl β -methylglutarate and octanoic acid in anhydrous methanol (10 A/dm², 2 A)³⁰⁸. The same authors synthesized the methyl ester of 10-methyloctadecanoic acid (38% yield) from the monomethyl ester of azelaic acid and 3-methylundecanoic acid (6.4–8.8 A/dm², 1.1–0.8 A); similarly, electrolysis of the monomethyl ester of γ -methyladipic acid with caproic acid gave 3-methyldecanoic acid (65% yield)¹⁶³; electrolytic condensation of a mixture of ethyl β , β -methylglutarate and capric or lauric acids³⁰⁹ yielded the corresponding β , β -methyltridecanoic and β , β -methylpentadecanoic acids.

Electrolysis in methanol at 10–20 A/dm² of monomethyl glutarate and 3-methylheptanoic acid, monomethyl succinate and 4-methyloctanoic acid, and the monomethyl ester of azelaic acid and 3,9-dimethyltridecanoic acid³¹⁰ yielded in the first two instances 6-methyldecanoic acid (35 and 32% yields), and in the last, 10,16-dimethyleicosanoic acid (54% yield). The same paper gives results on the "crossed" condensation when both initial acids contain branched carbon chains. Thus, 3,9-dimethyltridecanoic acid (51% yield) is formed in the electrolysis of monomethyl β -methylglutarate and 6-methyldecanoic acid. In an analogous manner 3,13,19-trimethyltridecanoic acid (18.5% yield) was prepared from monomethyl β -methylglutarate and 10,16-dimethyleicosanoic acid.

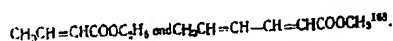
The electrolytic condensation of optically active substances has been very thoroughly investigated to find satisfactory methods for the synthesis of stereoisomers of compounds with branched chains of known structure. In this way it was possible to prepare D(+)-3-methylundecanoic acid (yield 35%)^{311–313} by electrolyzing an equimolar mixture of monomethyl L(+)- β -methylglutarate and caprylic acid in boiling methanol (Pt electrodes, 5 A, pH < 8)^{311, 313}.

By an analogous method L(-)-3-methylundecanoic acid (35.5% yield) was synthesized from monomethyl D(-)- β -methylglutarate and caprylic acid. Electrolysis of a mixture of the monomethyl ester of sebacic acid and D(+)-3-methylundecanoic or L(-)-3-methylundecanoic acid led to the formation of the corresponding D(+)-11-methylnonadecanoic acid (24% yield) and L(-)-11-methylnonadecanoic acid (22.5% yield). Condensation of palmitic acid with the monomethyl esters of D(+)- and L(-)- β -methylsuccinic acid gave the corresponding D(-)- and L(+)-2-methyloctadecanoic acids³¹⁴.

Higher yields can be obtained by using, in similar syntheses, an excess of the optically inactive component. In the electrolysis of caprylic acid with the monomethyl esters of L(+)- and D(-)- β -methylglutaric acid, in anhydrous methanol (molar ratio of the optically inactive to the active component 2:1, 2 A, 10 A/dm²), the corresponding D(+)- and L(-)-3-methylundecanoic acids (48% yield) were obtained. Under similar conditions, the monomethyl ester of azelaic acid and D(+)- and L(-)-3-methylundecanoic acid yielded the D(-)-10-methyloctadecanoic and L(+)-10-methyl-octadecanoic acids (30–31% yield)²¹⁵.

Unsaturated acids, mixed with dicarboxylic-acid esters, also undergo "crossed" electrolytic condensation. Thus pentadecenoic acid was obtained by the electrolysis of a mixture of undecenoic acid and monomethyl adipate in anhydrous methanol (0.83 A/dm², 80–100 V); after purification (via the methyl ester) the yield was approximately 15%.²¹⁶ In later investigations it was possible to increase the yield to 30%.¹⁷¹ A number of other unsaturated acids have also been synthesised by this reaction: 3,3-dimethyltetradec-13-enoic acid²¹⁷ by the electrolysis, in ethanol, of undecenoic acid and the ethyl ester of β , β -dimethylglutaric acid (10 A/dm², 100 V); docos-21-enoic acid from the ethyl ester of tetradec-1,14-dicarboxylic acid and oct-1-enoic acid²¹⁸.

When double bonds are present in the α -position to the carboxyl group unsatisfactory results are obtained; electrolysis of a mixture of acetic acid and monoethyl fumarate or monomethyl muconate in absolute methanol did not lead to the formation of the expected unsymmetrical products:



Bounds et al. showed that the geometrical configuration at the double bonds was maintained in the simple Kolbe reaction. They^{125,171,219} proved that the configuration at the double bonds was also preserved during the "crossed" condensation.

In the electrolysis of oleic and elaidic acids with an excess of adipic acid monomethyl ester, in anhydrous methanol (Pt electrodes, 10 A/dm², 1.7 A, 50°) the "crossed" products were obtained in both cases; their hydrolysis gave the corresponding erucic and brassidic acids (approximate yield 30%).¹⁷¹ Similarly, electrolysis of the monomethyl ester of suberic acid with oleic or elaidic acids²²⁰ gave, after hydrolysis, *cis*- and *trans*-tetracos-15-enoic acids respectively (respective yields 37% and 32%). The unsymmetrical product *cis*-octadec-11-enoic acid (12% yield), was obtained in the electrolysis of a mixture of *cis*-hexadec-9-enoic and monomethyl succinate²¹⁹.

Studies by American chemists²²⁰ on the preparation of esters of unsaturated α , ω -dicarboxylic acids by electrolysis a mixture of dicarboxylic-acid esters and conjugated dienes, are of great interest. They prepared the diethyl ester of deca-3,7-diene-1,10-dicarboxylic acid by the electrolysis, in methanol, of monoethyl malonate and butadiene (Pt electrodes, 1.1 A, -5°). The diethyl ester of dodeca-4,8-diene-1,12-dicarboxylic acid, the diethyl ester of hexadeca-6,10-diene-1,16-dicarboxylic acid, and other compounds were prepared by similar methods.

Satisfactory results have been obtained in the electrolytic condensation (in methanol) of acids of the acetylenic series²²¹. Stearolic acid (octadec-9-ynoic acid) (8% yield) was obtained from pentadec-6-ynoic acid and monomethyl glutarate (6–7 A/dm²). The same acid was obtained in the electrolytic condensation of tetradec-5-ynoic acid and

TABLE 11.

| In Mol reactants | | Molar ratio I:II | "Crossed" product | Yield, % (with respect to I) |
|------------------|---|------------------|---|------------------------------|
| I | II | | | |
| Levulinic acid | $\text{CH}_3(\text{CH}_2)_3\text{COOH}$ | 1 : 1 | methyl hexyl ketone | 35 |
| ditto | CH_3COOH | 1 : 3 | methyl propyl ketone | 20 |
| " | $\text{CH}_3(\text{CH}_2)_2\text{COOH}$ | 1 : 6 | methyl pentyl ketone | 62 |
| " | $\text{CH}_3\text{CH}_2\text{COOH}$ | 1 : 1 | methyl <i>n</i> -butyl ketone | 20 |
| " | $(\text{CH}_3)_2\text{CHCH}_2\text{COOH}$ | 1 : 1 | $\text{CH}_3\text{CO}(\text{CH}_2)_2\text{CH}(\text{CH}_3)_2$ | — |
| " | $\text{CH}_3(\text{CH}_2)_4\text{COOH}$ | 1 : 2 | methyl <i>n</i> -heptyl ketone | 42 |
| " | $\text{CH}_3(\text{CH}_2)_6\text{COOH}$ | 1 : 1 | methyl <i>n</i> -nonyl ketone | 31 |
| " | $\text{CH}_3(\text{CH}_2)_8\text{COOH}$ | 2 : 1 | methyl <i>n</i> -undecyl ketone | 24 |
| " | $\text{CH}_3(\text{CH}_2)_{10}\text{COOH}$ | 2 : 1 | $\text{CH}_3\text{CO}(\text{CH}_2)_8\text{CH}_3$ | 20 |
| " | $\text{CH}_3\text{OOC}(\text{CH}_2)_2\text{COOH}$ | 2 : 1 | $\text{CH}_3\text{CO}(\text{CH}_2)_4\text{COOCH}_3$ | 25–30 |
| " | $\text{CH}_3\text{OOC}(\text{CH}_2)_4\text{COOH}$ | 2 : 1 | $\text{CH}_3\text{CO}(\text{CH}_2)_6\text{COOCH}_3$ | 20 |
| " | $\text{CH}_3\text{OOC}(\text{CH}_2)_6\text{COOH}$ | 4 : 1 | $\text{CH}_3\text{CO}(\text{CH}_2)_8\text{COOCH}_3$ | 38 |
| " | $\text{CH}_3\text{OOC}(\text{CH}_2)_8\text{COOH}$ | 2 : 1 | $\text{CH}_3\text{CO}(\text{CH}_2)_{10}\text{COOCH}_3$ | 12 |

monomethyl adipate at 15–16 A/dm² (24% yield). Electrolysis of a mixture of monomethyl adipate and stearolic acid (8–10 A/dm²), pentanoic acid and the methyl ester of dodec-6-ynoic acid (6.8–8 A/dm²), and hexanoic acid and the methyl ester of dec-5-ynoic acid led to the formation of the corresponding docos-13-ynoic (behenolic) acid (26% yield), pentadec-6-ynoic acid (25% yield), and tetradec-5-ynoic acid (27% yield). Octadec-6-ynoic acid (23% yield) was prepared from the monomethyl ester of dec-5-yne-1,10-dicarboxylic acid and caprylic acid¹⁷⁸.

Halogeno-substituted acids were also subjected to this type of electrolytic condensation. Electrolysis of 10-fluorodecanoic and 5-chloropentanoic acid (Pt anode, 16 A/dm², 50°) in methanol gave 13-fluorotridecyl chloride (24.4% yield)¹²⁷. Under the same conditions 10-fluorodecanoic acid and the monomethyl ester of sebacic acid gave the methyl ester of 18-fluorostearic acid (4.8% yield).

The reactions of chloroacetic acid are of great interest; in the simple Kolbe reaction only oxidation products are obtained^{204–206}, but on electrolysis in a mixture with other acids its behaviour is similar to that of unsubstituted acetic acid. Cetyl chloride was obtained when chloroacetic and palmitic acids were electrolysed in a water-alcohol medium at 70°–75°.²²² Unsatisfactory results were obtained when trichloroacetic acid was used instead of chloroacetic acid.

Hunsdiecker²²³ did much work on the electrolytic condensation of keto-acids with fatty acids. He prepared dodeca-2,5-dione (30–45% yield) by the electrolysis of a mixture of 4,7-diketooctanoic and caproic acids in anhydrous methanol (3–5 A, 50–90 V). 4,7-Diketooctanoic acid was also reacted, under analogous conditions, with propionic, butyric, valeric, isovaleric, caprylic, and lauric acids, and also with monomethyl succinate and adipate. Unsymmetrical products were obtained in each case (no yields are quoted). He prepared dodeca-3,6-dione by electrolysis of a mixture of 4,7-diketooctanoic and valeric acids. Japanese investigators^{220,224,225} obtained very satisfactory results in the electrolytic condensation, in methanol, of keto-acids with saturated acids and half-esters of dicarboxylic acids (40 A/dm², 30°–35°) (Table 11).

Good results have been obtained with keto-acids and substituted aliphatic acids. Thus, 4,7-diketooctanoic acid was condensed with 5-methoxyvaleric and β -isopentyloxybutyric acid²²⁶; the yield of 9-isopentyloxydeca-2,5-dione was 30%.

The number of possible "crossed" electrolytic-condensation reactions of acids with various substituents is very large. The methyl ester of 14-methoxy-3-methyltetradecanoic acid (34.4% yield) has, for instance, been prepared by the electrolysis, in methanol, of monomethyl β -methylglutarate and 11-methoxyundecanoic acid (Ni anode, Pt cathode, 5-7 A, 80-100 V)²²⁶.

Hexadeca-2,15-dione (31% yield)²²⁷ was synthesised by electrolysis, in anhydrous methanol, of acetylvaleric and acetylparagonic acids (12.5 A/dm², 5 A, 50 V, 40°-50°).

The methyl ester of 10-acetamidodecanoic acid was obtained by electrolysis of 6-acetamidohexanoic acid and monomethyl adipate^{223,228}.

No "crossed" products were obtained in the electrolysis of monomethyl adipate with acetylglutamic and cyanoacetic acid (both substituted in the α -position to the carboxyl group)²²³. However, contrary to the general rule, Japanese investigators obtained satisfactory results in the simultaneous electrolytic condensation of methyl dodecylcyanoacetic acid and lauric acid, ethyl dodecylcyanoacetic acid and undecanoic acid²²⁹, ethyl dodecylcyanoacetic acid and capric acid, and also in the electrolysis of a mixture of the methyl ester of diisopentylmalonic acid and isocaproic acid²³⁰. (No reaction conditions or yields are given in these papers).

Hydrolysis of the products obtained in the electrolysis of monomethyl adipate with *threo*- and *erythro*-9,10-dihydroxystearic acid¹⁷¹ in anhydrous methanol, using Pt electrodes, yielded *threo*- and *erythro*-13,14-dihydroxybehenic acid (24% and 33% yields, respectively). *Threo*-11,12-dihydroxy arachidic acid was prepared by an analogous method¹²⁴ from monobenzyl succinate and *threo*-9,10-dihydroxystearic acid (13% yield).

Mixtures of aliphatic-aromatic and mixtures of aliphatic-aromatic and aliphatic acids have also been subjected to "crossed" electrolytic condensation. In one of the early investigations²¹⁰ the diethyl ester of α -methylhydrocinnamic acid was found in the electrolysis products of a mixture of acetic acid and monoethyl benzylmalonate. *t*-Pentylbenzene (32% yield) was obtained by electrolysis of 3-phenylisopentanoic acid in methanol, in the presence of an excess of sodium acetate (Pt anode, 15 A/dm², 60°)²⁷¹. Electrolysis of palmitic with phenylacetic and β -phenylpropionic acids led to the formation of the corresponding hexadecyl- and heptadecylbenzene¹⁷⁸. Electrolytic condensation of a mixture of γ -phenylbutyric acid with β -phenylpropionic or δ -phenylvaleric acids (Pt anode, 30°-40°) in a methanol-pyridine mixture (2:1)* gives, in the first case, small quantities of 1,5-diphenylpentane, and in the second 36% of 1,7-diphenylheptane²⁷⁰.

"Crossed" condensation as well as the simple Kolbe condensation can be carried out in alcoholic, water-alcohol, or aqueous media. Under these conditions (especially in the last case) side reactions occurring concurrently with the electrolytic condensation, decrease appreciably the yield of the desired product.

Japanese chemists investigated the electrolytic-condensation reaction in water-alcohol solutions. They electrolysed palmitic and acetic, propionic and butyric, and stearic and acetic acids²²² in dilute alcohol at 70°-75°, using a Pt anode, and obtained the expected products (no yields are

given). Electrolysis of the methyl ester of 2,12-dimethyltridecane-1,13-dicarboxylic acid with 6-methyldecanoic acid (Pt anode, 8 A/dm², pH 6.8-7, 30°) in 50% ethanol gave 3,13,19-trimethyltridecanoic acid (11% yield)^{331,332}. Under the same conditions the monoethyl ester of sebacic acid and 3,7-dimethyloctanoic acid³³³ gave the ethyl ester of 11,15-dimethylhexadecanoic acid. Stearic acid (21% yield)¹³⁸ was isolated from the electrolysis products of the monoethyl ester of sebacic acid and decanoic acid when the process was effected in a 35% aqueous-methanol solution at a molar ratio 1:3 of the initial components.

Electrolytic condensation of monoethyl adipate with some fatty acids in a water-alcohol solution³⁰⁴ gave good results (Table 12).

Greaves et al.¹³⁸ compared results obtained in the "crossed" electrolytic condensation in alcohol and water-alcohol solutions. Hexanoic acid (40% yield) was obtained by condensing monomethyl adipate with acetic acid in anhydrous methanol (1.5-2 A, 40°-50°) whilst in a water-alcohol solution (3 A, 30°) the yield was only 29% (the molar ratio of the ester to acetic acid was in both cases 1:6). Decanoic acid was prepared by the electrolysis of a mixture of monomethyl adipate and hexanoic acid (1:1); in anhydrous methanol the yield was 36%, in a water-alcohol medium 12%. The same acid was synthesised in 50% yield by electrolysis of the monomethyl ester of sebacic acid in aqueous methanol in the presence of a large excess of acetic acid (molar ratio of the starting materials 1:15).

The "crossed" synthesis in aqueous solutions has been investigated even less than in water-alcohol solutions; there are practically no published data.

In 1895³³⁴ the electrolytic condensation of monoethyl succinate with propionic, butyric, and isobutyric acids and of monoethyl malonate with acetic acid was carried out in an aqueous medium. The expected "crossed" products were obtained in each case.

Acetone¹⁹⁹ was prepared in 1900 by the electrolysis of an aqueous solution of potassium pyruvate and acetate in an electrolyser with a diaphragm (1.5-2 A, 17-18 V). The same author electrolysed in an aqueous medium (2.5 A and up to 20 V) mixtures of pyruvic and butyric acids, and also of levulinic and acetic or pyruvic acids, and obtained in the first two cases methyl propyl ketone, and in the last acetonylacetone.

The electrolysis of an aqueous solution of a mixture of the potassium salts of the monoethyl esters of malonic and suberic acids (1.7 A, 7.5 V, 15°) was carried out in 1922¹³⁴. After saponification of the electrolysis products, succinic and dodecane-1,2-dicarboxylic acids were obtained, as well

TABLE 12.

| Initial reactants | | Solvent | "Crossed" products (after hydrolysis) | Yield, % |
|--|--|--|--|----------|
| I | II | | | |
| H ₃ C ₁₀ COO(CH ₂) ₄ COOH | C ₆ H ₁₁ COOH | aqueous CH ₃ OH | CH ₃ (CH ₂) ₆ COOH | 23.3 |
| " | C ₇ H ₁₅ COOH | " | CH ₃ (CH ₂) ₁₀ COOH | 15 |
| " | C ₁₂ H ₂₅ COOH | aqueous C ₆ H ₅ OH | CH ₃ (CH ₂) ₁₄ COOH | 11.4 |
| " | (CH ₃) ₂ CHCH ₂ COOH | " | (CH ₃) ₂ CHCH ₂ (CH ₂) ₈ COOH | 8 |

* Pyridine is added to inhibit the formation of polymers which coat the anode during electrolysis.

as the "crossed" product - azelaic acid. The yields are not quoted, probably because they were small. Only Miller and Hofer³³¹ were able to record good results in similar electrolytic-condensation reactions. For example, using small quantities of reactants, they prepared ethyl butyrate (69.5% yield) by the electrolysis of the potassium salt of acetic acid and the ethyl ester of succinic acid.

Samokhvalov et al.³³² stated in 1952 that 3-5% yields of the "crossed" product can be obtained by electrolysis of 11-methoxyundecanoic acid and of monoethyl β -methylglutarate.

Recent investigations^{333,337} have shown that under certain conditions "crossed" electrolytic condensation can be successfully carried out in an aqueous medium. The electrolytic condensation of monoethyl adipate and 11-acetoxyundecanoic acid in an aqueous solution (24 A/dm², 40°-65°) can, after hydrolysis, give a 25-27% yield (based on 11-acetoxyundecanoic acid) of 15-hydroxypentadecanoic acid. Electrolysis of 11-hydroxyundecanoic acid and monoethyl adipate (molar ratio 1:3) gives a 7% yield of 15-hydroxypentadecanoic acid³³⁸.

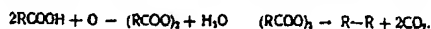
MECHANISM OF THE KOLBE REACTION

The mechanism of the Kolbe reaction is not entirely clear despite the large number of publications on the electrolytic condensation of carboxylic acids. The principal product and by-products have been investigated fairly thoroughly and in a large number of cases; several versions of the mechanism of transformation of the organic acids have been put forward. The nature of the anodic process, however - its kinetics, the role of the electrode surface in the condensation reaction, and the effect of the magnitude of the anode potential on the course of the reaction - have hitherto not been elucidated. In this respect, the electrolytic-condensation reaction does not appear to be an exception amongst other anodic processes including such apparently simple ones as, for example, the anodic evolution of oxygen from aqueous solutions.

Theories of the mechanism of electrolytic condensation developed in parallel with the extension of our knowledge of anodic processes and, in particular, of the evolution of oxygen.

In Kolbe's time it was assumed that only oxygen could be evolved at the anode during electrolysis of aqueous solutions; Kolbe, therefore, endeavoured to show that the electrolytic condensation was due to the oxidation of acids by oxygen, formed at the anode during the decomposition of water. The mechanism proposed by Kolbe was supported by several other investigators^{43,339}. These hypotheses, as well as Bourgoin's^{339,340} statement that anhydrides of the original acids are formed as intermediates in the Kolbe reaction, are only of historic interest.

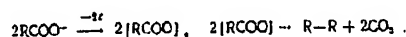
In 1896 Schall³⁴¹ suggested that carboxylic acids are oxidised by the active oxygen liberated at the anode, and that diacyl peroxides are formed; decomposition of the latter would give the product of the Kolbe synthesis:



Schall's hypothesis constituted the basis of Fichter's^{33,342-345} peroxide theory of electrolytic condensation which found many supporters^{105,346-348}.

According to Brown and Walker's theory⁶ the discharge of the carboxylic-acid anions and formation of uncharged radicals constitutes the primary process. The radicals

can be condensed with simultaneous splitting off of carbon dioxide:



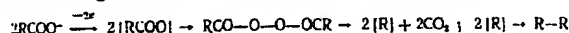
This was later developed as the free-radical theory^{15,16,91,103,104,108-110,349}.

Glasstone and Hickling's views, published in the 1930's^{8,37,45,51,350-353}, postulated that hydrogen peroxide, formed by discharge of OH^\cdot ions at the anode, is the basic intermediate product; this idea was closely associated with the mechanism of electrolytic condensation. The authors also used the results of their detailed investigations on the electrolytic condensation of acetic acid and malonic-acid half-ester^{6,37,45} in support of their theory, which is not, however, accepted by the majority of electrochemists at the present time.

The basic assumptions of the three theories of the electrolytic-condensation reaction are briefly discussed below.

The Peroxide Theory

This theory, proposed by Fichter in 1934³⁴⁴, starts from the following scheme of electrolytic condensation:



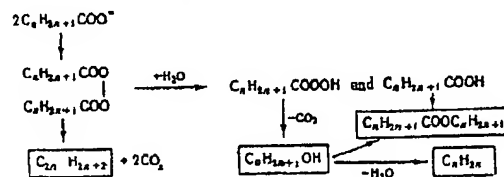
Unlike Schall, Fichter assumes that electrolytic condensation commences with the discharge of the carboxylic-acid anion, that diacyl peroxide is formed due to a further discharge of anions, and that the main product of the Kolbe synthesis is formed by the decomposition of the diacyl peroxide.

The formation of by-products is explained by the hydrolysis of the intermediate diacyl peroxide and the formation of the peracid which is converted to the alcohol by decarboxylation:



In an alkaline medium and at low concentration of the initial anions, this decomposition of the peroxide is the principal reaction and alcohol is the principal product (Hofer-Moest reaction).

The alcohols can then be dehydrated to olefins (with the same number of carbon atoms), or react with the free acid, or the discharged anion, to form esters. Electrolytic condensation of saturated fatty acids can be represented as follows (according to the peroxide theory):



This scheme does not take into account the formation of saturated hydrocarbons $\text{C}_n\text{H}_{2n+2}$, which are usually obtained in significant quantities.

Many attempts have been made to prove the formation of the diacyl peroxide at the anode. Positive results were obtained only when electrolysis was carried out under special conditions, and the yields of peroxides were very small. Fichter et al.^{129,354} succeeded in detecting diacyl peroxides and peracids during the electrolysis of caproate and the

half-ester of adipic acid, on strongly cooling the anode surface. Hallie¹¹⁴, who carried out electrolysis experiments at low temperatures, using a circulating electrolyte, proved that small quantities of diacyl peroxides are formed during the electrolysis of acetates and caproates.

Many papers mention that the composition of the products of electrolytic condensation and of the chemical decomposition of diacyl peroxides is similar, which is regarded as a proof of the peroxide theory by its supporters. Diacetyl peroxide, for instance, decomposes on heating to form ethane, methane, and carbon dioxide³⁶⁶⁻³⁶⁷.

A mixture of ethane and methane is formed when ultraviolet light acts on diacyl peroxides³⁶⁷. During the electrolysis of acetates, on the other hand, especially at low anodic current density, the formation of a mixture of ethane, methane, and carbon dioxide can also be observed^{18,89,99,368}.

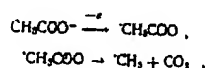
Fichter suggests that the peroxide theory is also confirmed by the formation of glycolaldehyde during the electrolysis of acetate³³ and monoethyl malonate⁷ in an ethylene-glycol solution. He regards the occurrence of the glycolaldehyde in both cases, as being due to the oxidation of ethylene glycol by the intermediate peroxides.

The Free-Radical Theory

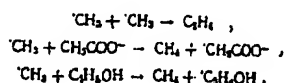
This theory is the one most thoroughly developed at the present time and has many supporters.

During investigations on side reactions in the electrolytic condensation Hölemann and Clusius^{103,104} carried out experiments on the electrolysis of deuterated acids dissolved in ordinary and in heavy water. They confirmed that a high current density is required in the electrolysis of acetates to produce ethane, and that a low current density (1 A/dm² and less) leads to the formation of methane. The deuterium content in the reaction products was proved to be determined solely by the composition of the acetate ion, not by the deuterium content of the water.

These results led the authors to the conclusion that the Kolbe reaction takes place via the formation of intermediate alkyl radicals – thus, in the electrolysis of acetates

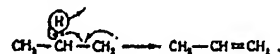


Dimerisation takes place readily at high current density because the concentration of the methyl radicals at the anode is high; at low current density the methyl radicals can react with the acetate ions or with the molecules of the organic solvent (e.g. with alcohol if this is added to the electrolyte) and form methane.

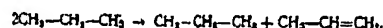


By following the isotope composition of the products in the electrolysis of deuteriopropionic^{15,109} and deuterobutyric¹¹⁰ acids, the authors concluded that unsaturated hydrocarbons (with half the number of carbon atoms of the main Kolbe-reaction products) are formed from the alkyl radicals, by the splitting off of the hydrogen atom originally in

the β -position to the carboxyl group. The hydrogen atoms of the molecule itself do not migrate.



The formation of propane as well as propylene in the electrolysis of butyric acid is explained by the disproportionation of hydrogen in the alkyl radicals:

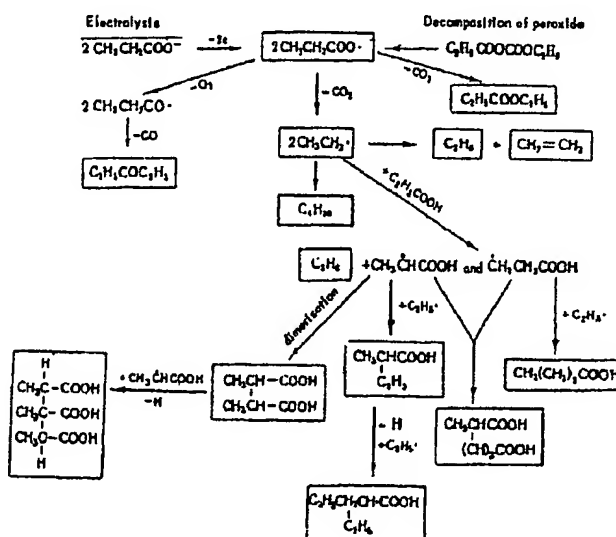


On electrolysis of an acid solution of α,α -dideuteriopropionic acid ($\text{CH}_3\text{CD}_2\text{COOH}$), Kruis and Schanzer¹⁵ obtained, among other products, a mixture of alcohols, containing 35% $\text{CH}_3\text{CD}_2\text{OH}$ and 65% $\text{CD}_3\text{CH}_2\text{OH}$, thus confirming that the oxidation of alkyl radicals to alcohols proceeds mainly at the β -position.

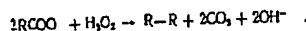
Goldschmidt and Stöckl¹⁰⁸ demonstrated the presence of free radicals during the electrolysis of fatty-acid salts by an indirect method, using the capacity of these radicals to initiate polymerisation of unsaturated compounds^{369,370} (styrene, acrylonitrile). The formation of alkylpyridines when pyridine is present in the electrolyte is also regarded as an indirect proof of the presence of radicals^{365,381}.

Goldschmidt et al. made a detailed study of the composition of the electrolysis products of propionates in anhydrous propionic acid and compared them with the decomposition products of dipropionyl peroxide¹⁶. Later investigations dealt with the decomposition of symmetrical and unsymmetrical peroxides of dicarboxylic acids³⁸². Goldschmidt stated that both the electrolysis of carboxylic acids and the decomposition of diacyl peroxides proceed via the formation of free radicals, but that the diacyl peroxides were not intermediate products in electrolysis.

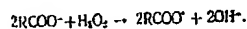
The electrolysis of propionates and decomposition of dipropionyl peroxides, according to Goldschmidt, is shown in the following scheme:



of aqueous solutions of fatty acids and form hydrogen peroxide; this reacts with the acid ions:



The formation of radicals represents an intermediate stage:



The latter may possibly combine and form diacyl peroxides – the intermediate products according to Fichter's theory:



When the concentration of hydrogen peroxide is not sufficiently high to effect the above reactions, or if the concentration of the anions is insufficient, hydrogen peroxide, or the active oxygen formed by its decomposition, will react with the acid and form the peracids:



The latter are decomposed to the alcohol and carbon dioxide:



Glasstone, Hickling, and Westwood carried out very detailed investigations on the electrolytic condensation of acetates^{37,45} and monoethyl malonate⁶ and endeavoured to correlate their results with the proposed mechanism, assuming the intermediate formation of hydrogen peroxide. The negative effect of the anode material (Ni, Fe, Au, etc.) and of foreign cations on the electrolytic condensation is related to their capacity to decompose hydrogen peroxide. The magnitude of the current density is related to the concentration of hydrogen peroxide set up in the layer close to the electrode. The effect of the pH of the medium on the yield of the main product is explained by the increased stability of hydrogen peroxide in acid solutions and by the decrease in the concentration of OH⁻ ions at lower pH.

Great importance is attached to the fact that the anode potential at which electrolytic condensation takes place is considerably higher than the potential at which oxygen is evolved from aqueous solutions. Hickling and Westwood⁶ state that during the electrolysis of a 2 M solution of monoethyl malonate: (a) the potential is about 2.8 V on smooth platinum (when good yields of the diethyl succinate are obtained), and (b) on platinised platinum, gold, and lead dioxide the potential is approximately 2.0 V, in consequence of which no electrolytic condensation takes place. The polarisation curve for the evolution of oxygen in a phthalate buffer with identical pH was obtained for comparison. Oxygen was evolved when the anode potential was approximately 2.0 V.

Commenting on Hickling and Glasstone's work⁴⁶, Nekrasov pointed out that their assumption that "hydrogen peroxide, formed at the anode, gives rise to the formation of radicals which are capable of establishing a much higher potential than the potential of oxygen evolution" cannot be upheld, since changes in the anode surface and increases in oxygen overvoltage can occur, without any participation of hydrogen peroxide. Indeed, many publications^{91,393} deny that the formation of hydrogen peroxide plays an indispensable part in anodic processes.

Glasstone and Hickling's proposed mechanism for the evolution of oxygen and other anodic processes is rejected

by most electrochemists. All attempts to effect the Kolbe synthesis by the action of hydrogen peroxide on acetate solutions have proved unsuccessful^{6,8,37}. The action of persulphates^{9,37,343,364,365} and of fluorine^{366,367}, however, gives the same products as the electrolytic condensation.

Obviously, the electrochemical aspects of the Kolbe reaction can be elucidated completely only when much more is known about anodic processes, especially about the mechanism of oxygen overvoltage at the anode.

At present it can only be stated that the evolution of oxygen at a platinum anode proceeds in two stages: (a) interaction of an OH⁻ ion with the metal electrode, forming a higher platinum oxide; (b) decomposition of the higher oxide with formation of oxygen.

Some authors^{368,369} assume that the discharge of OH⁻ ions is the slowest stage, determining the magnitude of the overvoltage. Others³⁷⁰⁻³⁷³ attach great importance to the decomposition of the higher platinum oxides, the formation of which has been demonstrated in a number of investigations³⁷¹⁻³⁷³.

It can be assumed that the discharge of the carboxylic-acid ion and its simultaneous interaction with the higher platinum oxides (formed at the electrode) is a necessary condition for the electrolytic condensation.

The anomalously large anode potentials during the reaction, recorded by Glasstone and Hickling, are not surprising as the presence of carboxylic-acid ions, especially at high current densities, can increase the oxygen overvoltage to a considerable degree; this also occurs during the anodic evolution of oxygen from solutions containing SO₃²⁻ ions³⁷⁴. Apparently, under these conditions the discharge of R-COO⁻ ions is energetically more probable than the discharge of hydroxyl ions, especially as the concentration of the latter is insignificant under conditions which are optimal for the electrolytic condensation.

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Research Institute for Synthetic
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THE MECHANISM OF TRANSPORT OF MATTER IN POROUS SORBENTS

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